

Study Guide

Key Concepts



14.1 Properties of Gases

- Gases are easily compressed because of the space between particles in a gas.
- The amount of gas (n), volume (V), and temperature (T) are factors that affect gas pressure (P).



14.2 The Gas Laws

- As the pressure of a gas increases, the volume decreases, if the temperature is constant.
- As the temperature of an enclosed gas increases, the volume increases, if the pressure is constant.
- As the temperature of an enclosed gas increases, the pressure increases, if the volume is constant.

- The combined gas law allows you to do calculations for situations in which only the amount of gas is constant.



14.3 Ideal Gases

- To calculate the number of moles of a contained gas requires an expression that contains the variable n .
- Real gases differ most from an ideal gas at low temperatures and high pressures.



14.4 Gases: Mixtures and Movements

- In a mixture of gases, the total pressure is the sum of the partial pressures of the gases.
- Gases of lower molar mass diffuse and effuse faster than gases of higher molar mass.

Vocabulary

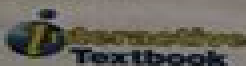
- Boyle's law (p. 418)
- Charles's law (p. 420)
- combined gas law (p. 424)
- compressibility (p. 413)
- Dalton's law of partial pressures (p. 432)
- diffusion (p. 435)
- effusion (p. 435)
- Gay-Lussac's law (p. 422)
- Graham's law of effusion (p. 435)
- ideal gas constant (p. 426)
- ideal gas law (p. 426)
- partial pressure (p. 432)

Key Equations

- Boyle's law: $P_1 \times V_1 = P_2 \times V_2$
- Charles's law: $\frac{V_1}{T_1} = \frac{V_2}{T_2}$
- Gay-Lussac's law: $\frac{P_1}{T_1} = \frac{P_2}{T_2}$
- Combined gas law: $\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2}$
- Ideal gas law: $P \times V = n \times R \times T$ or $PV = nRT$
- Dalton's law: $P_{\text{total}} = P_1 + P_2 + P_3 + \dots$
- Graham's law: $\frac{\text{Rate}_A}{\text{Rate}_B} = \sqrt{\frac{\text{molar mass}_B}{\text{molar mass}_A}}$

Organizing Information

Use these terms to construct a concept map that organizes the major ideas of this chapter.



Concept Map 14 Solve the Concept Map with the help of an interactive guided tutorial.

with **ChemASAP**

Boyle's law

Charles's law

partial pressure

combined gas law

diffusion

Prentice Hall Chemistry Section 14 Review Answers

Victor M. Corman

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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Applied Mechanics Reviews, 1987 **Journal of Industrial and Engineering Chemistry**, 1922 *The Journal of Industrial and Engineering Chemistry*, 1922 Encyclopedia of

Environmental Science and Engineering James R. Pfaflin, Edward N. Ziegler, 1992 First Published in 1992 Routledge is an imprint of Taylor Francis an informa company Thermodynamics of Solutions Eli Ruckenstein, Ivan L. Shulgin, 2009-06-17

This book consists of a number of papers regarding the thermodynamics and structure of multicomponent systems that we have published during the last decade Even though they involve different topics and different systems they have something in common which can be considered as the signature of the present book First these papers are concerned with difficult or very nonideal systems i e systems with very strong interactions e g hyd gen bonding between components or systems with large differences in the partial molar v umes of the components e g the aqueous solutions of proteins or systems that are far from normal conditions e g critical or near critical mixtures Second the conventional th modynamic methods are not sufficient for the accurate treatment of these mixtures Last but not least these systems are of interest for the pharmaceutical biomedical and related ind tries In order to meet the thermodynamic challenges involved in these complex mixtures we employed a variety of traditional methods but also new methods such as the fluctuation t ory of Kirkwood and Buff and ab initio quantum mechanical techniques The Kirkwood Buff KB theory is a rigorous formalism which is free of any of the proximations usually used in the thermodynamic treatment of multicomponent systems This theory appears to be very fruitful when applied to the above mentioned difficult systems

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