

Question	Answer	Marks	Guidance
5 (a)	F_2 forces mark F_2 has van der Waals' (forces) OR F_2 has induced dipole attractions OR interactions OR F_2 has temporary OR instantaneous dipole(-dipole) attraction OR interactions ✓ HCl forces mark HCl has permanent dipole(-dipole) attractions OR interactions ✓ Comparison of strength of forces between molecules mark intermolecular force in HCl is stronger than that in F_2 OR permanent dipoles are stronger (than induced dipoles) ✓ Boiling point mark more energy is required to break stronger (intermolecular) forces ✓	4	Use annotations ie ticks crosses ECF ^ etc for this part ALLOW vdWs for van der Waals' IGNORE F_2 has covalent bond for this mark IGNORE F_2 has 'intermolecular forces' <i>Quality of written communication: 'dipole(s)' spelled correctly and used in context for the second marking point</i> IGNORE HCl has 'intermolecular forces' IGNORE van der Waals' forces in HCl/ DO NOT ALLOW hydrogen bonding DO NOT ALLOW ionic bonding Look for strength of force comparison anywhere in the answer ALLOW ECF for hydrogen bonding in HCl being stronger than the stated intermolecular forces in F_2 BUT DO NOT ALLOW this mark if HCl or F_2 has covalent bonds broken OR if HCl has ionic bonds broken (the question asks for forces between molecules) IGNORE HCl has stronger van der Waals' (forces) than F_2 (as they both have the same number of electrons) DO NOT ALLOW fourth mark if covalent bonds are broken in HCl or F_2 OR if ionic bonds are broken in HCl IGNORE 'heat' but ALLOW 'heat energy'

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Table of Contents Ocr F321 Jan 2013 Paper

1. Understanding the eBook Ocr F321 Jan 2013 Paper
 - The Rise of Digital Reading Ocr F321 Jan 2013 Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr F321 Jan 2013 Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr F321 Jan 2013 Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr F321 Jan 2013 Paper
 - Personalized Recommendations
 - Ocr F321 Jan 2013 Paper User Reviews and Ratings
 - Ocr F321 Jan 2013 Paper and Bestseller Lists
5. Accessing Ocr F321 Jan 2013 Paper Free and Paid eBooks
 - Ocr F321 Jan 2013 Paper Public Domain eBooks
 - Ocr F321 Jan 2013 Paper eBook Subscription Services

- Ocr F321 Jan 2013 Paper Budget-Friendly Options
- 6. Navigating Ocr F321 Jan 2013 Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr F321 Jan 2013 Paper Compatibility with Devices
 - Ocr F321 Jan 2013 Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr F321 Jan 2013 Paper
 - Highlighting and Note-Taking Ocr F321 Jan 2013 Paper
 - Interactive Elements Ocr F321 Jan 2013 Paper
- 8. Staying Engaged with Ocr F321 Jan 2013 Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr F321 Jan 2013 Paper
- 9. Balancing eBooks and Physical Books Ocr F321 Jan 2013 Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr F321 Jan 2013 Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr F321 Jan 2013 Paper
 - Setting Reading Goals Ocr F321 Jan 2013 Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr F321 Jan 2013 Paper
 - Fact-Checking eBook Content of Ocr F321 Jan 2013 Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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