

- (a) N_2 (b) O_2
 (c) Cl_2 (d) S

10. Which gas is used to manufacture king of chemicals sulphuric acid?

- (a) N_2 (b) O_2
 (c) Cl_2 (d) S

11. Equilibrium constant has no unit when number of moles of reactants and products are;

- (a) same (b) different
 (c) both a & b (d) none

12. For reactions having large K_c value, the reaction proceeds to;

- (a) completion
 (b) equilibrium state
 (c) back ward (d) None

13. The characteristics of reversible reactions are the following except;

- (a) products never recombine to form reactants
 (b) they never complete
 (c) they proceed in both ways
 (d) they have a double arrow between reactants and products

14. In the lime kiln, the reaction



goes to completion because;

- (a) of high temperature
 (b) CaO is more stable than $CaCO_3$
 (c) CO_2 escapes continuously
 (d) CaO is not dissociated

15. For the reaction,



the expression for the equilibrium constant is:

- (a) $\frac{[2A][B]}{[3C]}$ (b) $\frac{[A]^2[B]}{[C]^3}$

- (c) $\frac{[3C]}{[2A][B]}$ (d) $\frac{[C]^3}{[A]^2[B]}$

16. When a system is at equilibrium states?

- (a) the concentration of reactants and products becomes equal
 (b) the opposing reactions (forward and reverse) stop
 (c) the rate of the reverse reaction becomes very low
 (d) the rates of the forward and reverse reactions becomes equal.

17. Which one of the following statements is not correct about active mass?

- (a) rate of reaction is directly proportional to active mass.
 (b) active mass is taken in molar concentration
 (c) active mass is represented by square brackets
 (d) active mass means total mass of substances

18. When the magnitude of K_c is very large it indicates;

- (a) reaction mixture consists of almost all products
 (b) reaction mixture has almost all reactants
 (c) reaction has not gone to completion
 (d) reaction mixture has negligible products

19. When the magnitude of K_c is very small it indicates;

- (a) equilibrium will never establish

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1. Understanding the eBook Objective Type Question Of Chemistry For 10 Th
 - The Rise of Digital Reading Objective Type Question Of Chemistry For 10 Th
 - Advantages of eBooks Over Traditional Books
2. Identifying Objective Type Question Of Chemistry For 10 Th
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Objective Type Question Of Chemistry For 10 Th
 - User-Friendly Interface

4. Exploring eBook Recommendations from Objective Type Question Of Chemistry For 10 Th
 - Personalized Recommendations
 - Objective Type Question Of Chemistry For 10 Th User Reviews and Ratings
 - Objective Type Question Of Chemistry For 10 Th and Bestseller Lists
5. Accessing Objective Type Question Of Chemistry For 10 Th Free and Paid eBooks
 - Objective Type Question Of Chemistry For 10 Th Public Domain eBooks
 - Objective Type Question Of Chemistry For 10 Th eBook Subscription Services
 - Objective Type Question Of Chemistry For 10 Th Budget-Friendly Options
6. Navigating Objective Type Question Of Chemistry For 10 Th eBook Formats
 - ePub, PDF, MOBI, and More
 - Objective Type Question Of Chemistry For 10 Th Compatibility with Devices
 - Objective Type Question Of Chemistry For 10 Th Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Objective Type Question Of Chemistry For 10 Th
 - Highlighting and Note-Taking Objective Type Question Of Chemistry For 10 Th
 - Interactive Elements Objective Type Question Of Chemistry For 10 Th
8. Staying Engaged with Objective Type Question Of Chemistry For 10 Th
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Objective Type Question Of Chemistry For 10 Th
9. Balancing eBooks and Physical Books Objective Type Question Of Chemistry For 10 Th
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Objective Type Question Of Chemistry For 10 Th
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Objective Type Question Of Chemistry For 10 Th
 - Setting Reading Goals Objective Type Question Of Chemistry For 10 Th
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Objective Type Question Of Chemistry For 10 Th
 - Fact-Checking eBook Content of Objective Type Question Of Chemistry For 10 Th
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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