

# Physics

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(Chapter 11) (Thermal Properties of Matter) (Exercises)  
(Class 11)

$$\text{Useful Energy, } \Delta Q = \frac{50}{100} \times 1.5 \times 10^6 = 7.5 \times 10^5 \text{ J}$$

$$\text{But } \Delta Q = mc\Delta T$$

$$\therefore \Delta T = \frac{\Delta Q}{mc} = \frac{7.5 \times 10^5}{8 \times 10^3 \times 0.91} = 103^\circ\text{C}$$

Therefore, in 2.5 minutes of drilling, the rise in the temperature of the block is  $103^\circ\text{C}$ .

## Question 11.13:

A copper block of mass 2.5 kg is heated in a furnace to a temperature of  $500^\circ\text{C}$  and then placed on a large ice block. What is the maximum amount of ice that can melt? (Specific heat of copper =  $0.39 \text{ J g}^{-1} \text{ K}^{-1}$ ; heat of fusion of water =  $335 \text{ J g}^{-1}$ )

## Answer 11.13:

Mass of the copper block,  $m = 2.5 \text{ kg} = 2500 \text{ g}$

Rise in the temperature of the copper block,  $\Delta\theta = 500^\circ\text{C}$

Specific heat of copper,  $C = 0.39 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$

Heat of fusion of water,  $L = 335 \text{ J g}^{-1}$

The maximum heat the copper block can lose,  $Q = mC\Delta\theta = 2500 \times 0.39 \times 500 = 487500 \text{ J}$

Let  $m_1 \text{ g}$  be the amount of ice that melts when the copper block is placed on the ice block.

The heat gained by the melted ice,  $Q = m_1 L$

$$\therefore m_1 = \frac{Q}{L} = \frac{487500}{335} = 1455.22 \text{ g}$$

Hence, the maximum amount of ice that can melt is 1.45 kg.

## Question 11.14:

In an experiment on the specific heat of a metal, a 0.20 kg block of the metal at  $150^\circ\text{C}$  is dropped in a copper calorimeter (of water equivalent 0.025 kg) containing  $150 \text{ cm}^3$  of water at  $27^\circ\text{C}$ . The final temperature is  $40^\circ\text{C}$ . Compute the specific heat of the metal. If heat losses to the surroundings are not negligible, is your answer greater or smaller than the actual value for specific heat of the metal?

## Answer 11.14:

Mass of the metal,  $m = 0.20 \text{ kg} = 200 \text{ g}$

Initial temperature of the metal,  $T_1 = 150^\circ\text{C}$  and Final temperature of the metal,  $T_2 = 40^\circ\text{C}$

Calorimeter has water equivalent of mass,  $m' = 0.025 \text{ kg} = 25 \text{ g}$

Volume of water,  $V = 150 \text{ cm}^3$

Mass ( $M$ ) of water at temperature  $T = 27^\circ\text{C}$ :  $150 \times 1 = 150 \text{ g}$

Fall in the temperature of the metal:  $\Delta T = T_1 - T_2 = 150 - 40 = 110^\circ\text{C}$

Specific heat of water,  $C_w = 4.186 \text{ J/g/}^\circ\text{K}$  and Specific heat of the metal =  $C$

Heat lost by the metal,  $\theta = mC\Delta T \dots (i)$

Rise in the temperature of the water and calorimeter system:

$$\Delta T' = 40 - 27 = 13^\circ\text{C}$$

Heat gained by the water and calorimeter system:  $\Delta\theta' = m_1 C_w \Delta T' = (M + m') C_w \Delta T' \dots (ii)$

Heat lost by the metal = Heat gained by the water and calorimeter system

$$mC\Delta T = (M + m') C_w \Delta T' \Rightarrow 200 \times C \times 110 = (150 + 25) \times 4.186 \times 13$$

$$\therefore C = \frac{175 \times 4.186 \times 13}{110 \times 200} = 0.43 \text{ J g}^{-1} \text{ K}^{-1}$$

If some heat is lost to the surroundings, then the value of  $C$  will be smaller than the actual value.

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### **Table of Contents Physics Class 11 Chapter Wise Question**

1. Understanding the eBook Physics Class 11 Chapter Wise Question
  - The Rise of Digital Reading Physics Class 11 Chapter Wise Question
  - Advantages of eBooks Over Traditional Books
2. Identifying Physics Class 11 Chapter Wise Question
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physics Class 11 Chapter Wise Question
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physics Class 11 Chapter Wise Question
  - Personalized Recommendations

- Physics Class 11 Chapter Wise Question User Reviews and Ratings
- Physics Class 11 Chapter Wise Question and Bestseller Lists
- 5. Accessing Physics Class 11 Chapter Wise Question Free and Paid eBooks
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  - Physics Class 11 Chapter Wise Question eBook Subscription Services
  - Physics Class 11 Chapter Wise Question Budget-Friendly Options
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  - Physics Class 11 Chapter Wise Question Compatibility with Devices
  - Physics Class 11 Chapter Wise Question Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physics Class 11 Chapter Wise Question
  - Highlighting and Note-Taking Physics Class 11 Chapter Wise Question
  - Interactive Elements Physics Class 11 Chapter Wise Question
- 8. Staying Engaged with Physics Class 11 Chapter Wise Question
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physics Class 11 Chapter Wise Question
- 9. Balancing eBooks and Physical Books Physics Class 11 Chapter Wise Question
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physics Class 11 Chapter Wise Question
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physics Class 11 Chapter Wise Question
  - Setting Reading Goals Physics Class 11 Chapter Wise Question
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physics Class 11 Chapter Wise Question
  - Fact-Checking eBook Content of Physics Class 11 Chapter Wise Question

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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