

16

Electrostatics

16-1 Electrostatic Force

Vocabulary

Electrostatics: The study of electric charges, forces, and fields.

The symbol for electric charge is the letter “ q ” and the SI unit for charge is the **coulomb (C)**. The coulomb is a very large unit.

$$1 \text{ C} = 6.25 \times 10^{18} \text{ electrons} \quad \text{or} \\ 1 \text{ electron has a charge of } 1.60 \times 10^{-19} \text{ C.}$$

Electrons surrounding the nucleus of an atom carry a negative charge. Protons, found inside the nucleus of the atom, carry a positive charge of $1.60 \times 10^{-19} \text{ C}$, while neutrons (which also reside in the nucleus) are neutral. It is important to remember that only electrons are free to move in a substance. Protons and neutrons usually do not move.

When two objects with like charges, positive or negative, are brought near each other, they experience a repulsive force. When objects with opposite charges, one negative and one positive, are brought side by side, they experience an attractive force. These forces can be described with Coulomb’s law.

Vocabulary

Coulomb’s Law: Two charged objects attract each other with a force that is proportional to the charge on the objects and inversely proportional to the square of the distance between them.

$$F \propto \frac{q_1 q_2}{d^2}$$

This equation looks very similar to Newton’s law of universal gravitation. As before, the sign $\propto$ means “proportional to.” To make an equation out of this proportionality, insert a quantity called the **electrostatic constant, k** .

$$k = 9.0 \times 10^9 \text{ N} \cdot \text{m}^2/\text{C}^2$$

The magnitude of Coulomb’s law can now be written as an equation.

$$\text{electrostatic force} = \frac{(\text{electrostatic constant})(\text{charge 1})(\text{charge 2})}{(\text{distance})^2} \quad \text{or} \quad F = \frac{kq_1 q_2}{d^2}$$

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Léo Pomar



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