

**QUESTION 2**

- 2.1 The graph of molecular size versus the boiling point is given below.  
The letters **A**, **B** and **C** represent the compounds  $\text{CH}_4$ ,  $\text{C}_2\text{H}_6$  and  $\text{C}_3\text{H}_8$  respectively.




- 2.1.1 Define the term *boiling point*. (2)
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)
- 2.2 Consider the two molecules in the table below.

NAME OF SUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT ( $^{\circ}\text{C}$ )
Ammonia	$\text{NH}_3$	17	- 33
Phosphine	$\text{PH}_3$	x	- 87,4

- 2.2.1 Determine the molecular mass (x) of phosphine (1)
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

# Physical Sciences Grade11 Caps Question Papers

**Karin H. Kelder**



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web equilibrium questions a covalent bond has a standard gibbs free energy of  $400 \text{ kJ mol}^{-1}$  while the hydrogen bond and van der waals forces have substantially lower values respectively  $25$  and  $5 \text{ kJ mol}^{-1}$  given such  $\Delta G$  values which of the following is equivalent to the ratio of the  $K_{eq}$  value of a covalent bond to that of a hydrogen bond

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web jun 16 2020 summary skills to develop to understand what is meant by chemical equilibrium two methods to describe the likely result of a chemical reaction in the previous chapter we discussed the principles of chemical kinetics which deal with the rate of change or how quickly a given chemical reaction occurs

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web nov 13 2022 define the equilibrium state of a chemical reaction system what is its practical significance state the meaning and significance of the following terms reversible reaction quantitative reaction kinetically inhibited reaction explain the meaning of the statement equilibrium is macroscopically static but microscopically dynamic very

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web in a chemical equilibrium the forward and reverse reactions occur at equal rates and the concentrations of products and reactants remain constant if we run a reaction in a closed system so that the products cannot escape we often find the reaction does not give a 100 yield of products