

Molarity Calculations

Calculate the molarities of the following solutions:



- 1) 2.3 moles of sodium chloride in 0.45 liters of solution.

$$M = \frac{2.3 \text{ moles}}{0.45 \text{ L}} = 5.111 \frac{\text{mol}}{\text{L}} = \boxed{5.11 \text{ M}}$$



- 2) 1.2 moles of calcium carbonate in 1.22 liters of solution.

$$M = \frac{1.2 \text{ mol}}{1.22 \text{ L}} = 0.983606557 \frac{\text{mol}}{\text{L}} = \boxed{0.98 \text{ M}}$$



- 3) 0.09 moles of sodium sulfate in 12 mL of solution.

12 mL = 0.012 L

$$M = \frac{0.09 \text{ mol}}{0.012 \text{ L}} = \boxed{7.5 \text{ M}}$$



- 4) 0.75 moles of lithium fluoride in 65 mL of solution.

65 mL = 0.065 L

$$M = \frac{0.75 \text{ mol}}{0.065 \text{ L}} = 11.53846154 \text{ M} = \boxed{11.54 \text{ M}}$$



- 5) 0.8 moles of magnesium acetate in 5 liters of solution.

$$M = \frac{0.8 \text{ mol}}{5 \text{ L}} = \boxed{0.16 \text{ M}}$$



- 6) 120 grams of calcium nitrite in 240 mL of solution.

240 mL = 0.240 L

$$n = 120 \text{ g} \left(\frac{1 \text{ mol}}{132.09 \text{ g}} \right)$$

MW_{Ca(NO₂)₂} = 132.09 g/mol

n = 0.91 mol

$$M = \frac{0.91 \text{ mol}}{0.240 \text{ L}} = \boxed{3.79 \text{ M}}$$



- 7) 98 grams of sodium hydroxide in 2.2 liters of solution.

MW_{NaOH} = 39.997 g/mol

$$n = 98 \text{ g} \left(\frac{1 \text{ mol}}{39.997 \text{ g}} \right) = 2.45 \text{ mol}$$

$$M = \frac{2.45 \text{ mol}}{2.2 \text{ L}} = \boxed{1.11 \text{ M}}$$



- 8) 1.2 grams of hydrochloric acid in 25 mL of solution.

25 mL = 0.025 L

MW_{HCl} = 36.461 g/mol

$$n = (1.2 \text{ g}) \left(\frac{1 \text{ mol}}{36.461 \text{ g}} \right) = 0.033 \text{ mol}$$

$$M = \frac{0.033 \text{ mol}}{0.025 \text{ L}} = \boxed{1.32 \text{ M}}$$



- 9) 45 grams of ammonia in 0.75 L of solution.

MW_{NH₃} = 17.031 g/mol

$$n = 45 \text{ g} \left(\frac{1 \text{ mol}}{17.031 \text{ g}} \right) = 2.64 \text{ mol}$$

$$M = \frac{2.64 \text{ mol}}{0.75 \text{ L}} = \boxed{3.52 \text{ M}}$$

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DJ Losen



Page 3 Chemfiesta More Molecular Problems:

Page 3 Chemfiesta More Molecular Problems Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Page 3 Chemfiesta More Molecular Problems**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we shall delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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Table of Contents Page 3 Chemfiesta More Molecular Problems

1. Understanding the eBook Page 3 Chemfiesta More Molecular Problems
 - The Rise of Digital Reading Page 3 Chemfiesta More Molecular Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Page 3 Chemfiesta More Molecular Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Page 3 Chemfiesta More Molecular Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Page 3 Chemfiesta More Molecular Problems
 - Personalized Recommendations
 - Page 3 Chemfiesta More Molecular Problems User Reviews and Ratings
 - Page 3 Chemfiesta More Molecular Problems and Bestseller Lists

5. Accessing Page 3 Chemfiesta More Molecular Problems Free and Paid eBooks
 - Page 3 Chemfiesta More Molecular Problems Public Domain eBooks
 - Page 3 Chemfiesta More Molecular Problems eBook Subscription Services
 - Page 3 Chemfiesta More Molecular Problems Budget-Friendly Options
6. Navigating Page 3 Chemfiesta More Molecular Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Page 3 Chemfiesta More Molecular Problems Compatibility with Devices
 - Page 3 Chemfiesta More Molecular Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Page 3 Chemfiesta More Molecular Problems
 - Highlighting and Note-Taking Page 3 Chemfiesta More Molecular Problems
 - Interactive Elements Page 3 Chemfiesta More Molecular Problems
8. Staying Engaged with Page 3 Chemfiesta More Molecular Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Page 3 Chemfiesta More Molecular Problems
9. Balancing eBooks and Physical Books Page 3 Chemfiesta More Molecular Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Page 3 Chemfiesta More Molecular Problems
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Page 3 Chemfiesta More Molecular Problems
 - Setting Reading Goals Page 3 Chemfiesta More Molecular Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Page 3 Chemfiesta More Molecular Problems
 - Fact-Checking eBook Content of Page 3 Chemfiesta More Molecular Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

Page 3 Chemfiesta More Molecular Problems Introduction

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