

Name: _____

Date: _____

Chemistry Physics NJASK Practice Test

1. The data table displays the mass and volume of three different objects.

	Object A	Object B	Object C
Mass	12 g	6 g	5 g
Volume	6 cm ³	3 cm ³	5 cm ³

Using the formula: $\text{Density} = \text{Mass}/\text{Volume}$

Which of the following statements is correct?

- A. Object B is more dense than object A.
 - B. Objects A and B have equal densities.
 - C. Object C is more dense than Object A.
 - D. Objects B and C have equal densities.
2. Atoms are composed of protons, neutrons and electrons. What is the electric charge of the neutron?
- A. Positive
 - B. Negative
 - C. Neutral
 - D. Double-charged
3. A battery is placed in a flashlight. The battery is attached to wires that connect to the lightbulb. Which of the answers below shows the correct order of the energy changes?
- A. Chemical → Electricity → Light
 - B. Chemical → Mechanical → Light
 - C. Light → Electrical → Chemical
 - D. Nuclear → Mechanical → Light
4. Which of the following is a chemical reaction?
- A. A nail rusting
 - B. Water freezing
 - C. Sugar dissolving in water
 - D. Glass shattering

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ/mole}$

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