#### CHEMISTRY 1211 TEST #3

Use a Scantron sheet for Multiple Choices (3pts each). Do not write on the Periodic Table.

Use back of this page to show your work/answer for the Big Question.

1mi=1.609km; 1lbs=453.6g; 1gal=3.785L; 1mL=1cm³; 1=10 <sup>100</sup>; m; 1cal=4.184J; g=9.807m/s 1atm=760mmHg; Density(Hg)=13.60g/mL; c=2.9979x10 <sup>100</sup>m/s; h=6.626x10 <sup>100</sup>J R<sub>0</sub>=2.180x10 s; <sup>100</sup>J; m<sub>0</sub>=9.109x10 <sup>100</sup>kg; R=8.315J/kmol.

Which photon has the lowest energy?

A( X-ray), B( Ultarviolet radiation, UV), C(Visible light), D(Infrared radiation, IR), E(Microwave).

- Energy of an electromagnetic wave increases with (a) increasing, (b) decreasing, frequency; therefore a blue light has a higher energy than a red light because of its (c) higher, (d) lower frequency than that of a red light. A(a,c), B(a,d), C(b,c), D(b,d).
- Frequency of a radio wave is 104.7 MHz. It's wavelength must be about A(3 †m), B(3 cm), C(3 m), D(3 km), E(3 Mm).
- Find energy of an X-ray with a wavelength of 1.0x10 to m (= 0.10 nm). Report it (a) per photon and (b) per mole.

A(1.986x10<sup>-12</sup>/photon; 1.196x10<sup>-1</sup>kJ/mol), B(1.986x10 --- <sup>12</sup>J/photon; 1.196x10<sup>-1</sup>kJ/mol), C(1.986x10<sup>-1</sup>J/photon; 1.196x10<sup>-1</sup>kJ/mol), D(1.986x10 --- <sup>12</sup>J/photon; 1.196x10<sup>-1</sup>kJ/mol).

- Who got the Nobel Prize in 1921 with an idea that light behave as if it is a particle whose energy is directly proportional to its frequency?
   A(Balmer), B(Planck), C(Einstein), D(Bohr), E(Heisenberg).
- 6. What is the energy (J) of an electron at the third energy level (i.e., in the L shell with the principle quantum number n=3) in a H atom?

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