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# Printable Notary Journal

**J Spring**



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web equilibrium moles  $HSCN$   $2.00 \times 10^{-5} \text{ mol}$   $3.00 \times 10^{-6} \text{ mol}$   $1.70 \times 10^{-5} \text{ mol}$   $HSCN$  knowing the number of moles of  $Fe^{3+}$  and

hscn present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species  
 $\text{mol Fe}^{3+} = 1.70 \times 10^{-5} \text{ mol}$   
 $\text{mol Fe}^{3+} = 8.50 \times 10^{-4} \text{ mol}$

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