

Marlene Thelenberg-Klarwein - 27. November 2022

Skizzen:

Polypus, Monoclonales Antikörper, Antikörperproduktion, Antikörperreinigung, Antikörper

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- Antikörper ist auch Antikörper für Antigen, Antikörper
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# Physics Fundamentals Note Taking Guide Gpb Episode 13answers

**Yicheng Fang**



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fraction  $x$  is the ratio of moles of one substance in a mixture to the total number of moles of all substances for a mixture of

two substances  $c_e a$  and  $c_e b$  the mole fractions of each would be written as

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water  $n_a$  54 g 18 g mol 1 3 mol total number of moles  $n_a + n_b$  0 5 3 3 5 mol mole fraction of solute ethyl alcohol  $x_b = \frac{n_b}{n_a + n_b}$

0 5 3 5 0 1429

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40 g of  $H_2O$  molar mass of  $H_2O$  is 18gm mole and molar mass of  $CH_3OH$  is 32gm mole solution moles of  $CH_3OH$  4 5 32 0 14

mole moles of  $H_2O$  40 18 2 2 moles therefore according to the equation

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web we used mole fractions in gen chem 1 to describe gas mixtures where the partial pressure of a gas was its mole fraction

time the total pressure  $p_a \times a \times t$  mole fraction  $a$  left frac mole  $a$  moles total right sum  $n_1$  all  $x_n$  1 mole fraction of substance



a is often denoted by  $x_a$  and

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web 1 in the given mixture the sum of all the mole fractions is equal to one  $x_a + x_b + \dots = 1$  2 the multiplication of the mole fraction

by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction

mole fraction is not dependent on the temperature

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web nov 6 2023  $x_a$  mol a mol a mol b and  $x_b$  mol b mol a mol b note in the given mixture the sum of all the mole fractions

present is equal to one  $x_a + x_b + \dots = 1$  when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is

a unitless and dimensionless expression image will be uploaded soon

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ways to represent the concentration of the solution it is defined as the number of molecules of component a divided by the

total number of molecules of all the components in the solution

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web solution total mass of the answer 100 g mass of benzene 30 g therefore mass of carbon tet 100 30 g benzene s molar

mass  $C_6H_6$  6 12 6 1 g mol 1 78 g mol 1 hence  $C_6H_6$  s number of moles 30 78 mol 0 3846 mol molar mass of carbon tet  $CCl_4$  1

12 4 355 154 g mol 1

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web jan 30 2020 ans the molality of solution is 1 mol kg 1 or 1 m example 02 11 11 g of urea  $NH_2CONH_2$  was dissolved in 100

g of water calculate the molarity and molality of the solution given n 14 h 1 c 12 o 16 given mass of solute urea 11 11 g mass

of solvent water 100 g 0 1 kg to find molarity of solution solution

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal kbr that

contains 2 92 kg of water 1 molality moles solute divided by kilograms solute 0 449 mol kg x 2 92 kg x 1 31108 mol of kbr 2

moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g kbr

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web calculate the mass percent mole fraction molality and molarity of hexane c 6 h 14 d 0 660 g cm 3 when it is mixed with heptane c 7 h 12 d 0 684 g cm3 in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web since the mole fraction of the  $O_2$  is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.210 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of  $O_2$  are mixed in a flask at room temperature

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