

9.1

THE ARITHMETIC OF EQUATIONS

SECTION REVIEW

Objectives

- Calculate the amount of reactants required or product formed in a chemical process.
- Interpret balanced chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

Key Terms

- stoichiometry

Part A Completion

Copy this completion exercise for class. Your understanding of the concepts and terms is clear when indicated in this exercise. Blank blanks are the completed work of others; absent phrases are none of these.

- The coefficients in a balanced chemical equation indicate _____ 1. _____ the relative number of _____ 2. _____ of reactants and products. All stoichiometric calculations begin with a _____ 3. _____ On a _____ 4. _____ and _____ 5. _____ are concerned in every reaction: moles, volumes, and representative particles may not be. _____ 6. _____
- In solving stoichiometric problems, conversion factors relating moles of reactants to _____ 7. _____ of products are used. _____ 8. _____ If you assume _____ 9. _____, the equation also tells you about the volume of gas.

Part B True-False

(Quickly mark off your answers to the statements as T for true or F for false; some answers may be C for cannot tell, or D for don't know.)

- _____ 1. The coefficients in a balanced chemical equation can be used to determine the relative limiting reactants to products.
- _____ 2. The coefficients in a balanced chemical equation tell the relative volumes of reactants and products, expressed in any suitable units of volume.
- _____ 3. To calculate the mass of a molecule in grams, you can use the molar mass and Avogadro's number.

Section 9 The Arithmetic Of Equations Answers

Michael Smith



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