

Chapter 6 Chemical Bonds

Section 6.2 Covalent Bonding**(pages 165–169)**

This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules.

Reading Strategy (page 165)

Relating Text and Visuals As you read the section, look closely at Figure 9. Complete the table by describing each type of model shown. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Molecular Models	
Model	Description
Electron dot	
Structural formula	
Space-filling	
Electron cloud	

Covalent Bonds (pages 165–167)

- Describe a covalent bond. _____

- Circle the letters of molecular models that show orbitals of atoms overlapping when a covalent bond forms.
 - electron dot
 - structural formula
 - space-filling
 - electron cloud
- Describe a molecule. _____

- Is the following sentence true or false? In a covalent bond, the atoms are held together by the attractions between the shared electrons and the protons in each nucleus. _____
- Circle the correct answer. Nitrogen has five valence electrons. How many pairs of electrons must two nitrogen atoms share in order for each atom to have eight valence electrons?
 - zero
 - one
 - two
 - three

Section 8 2 The Nature Of Covalent Bonding Guided Reading

**Thirumalaisamy P.
Velavan**corresponding



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Table of Contents Section8 2 The Nature Of Covalent Bonding Guided Reading

1. Understanding the eBook Section8 2 The Nature Of Covalent Bonding Guided Reading
 - The Rise of Digital Reading Section8 2 The Nature Of Covalent Bonding Guided Reading
 - Advantages of eBooks Over Traditional Books
2. Identifying Section8 2 The Nature Of Covalent Bonding Guided Reading
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section8 2 The Nature Of Covalent Bonding Guided Reading
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section8 2 The Nature Of Covalent Bonding Guided Reading
 - Personalized Recommendations
 - Section8 2 The Nature Of Covalent Bonding Guided Reading User Reviews and Ratings
 - Section8 2 The Nature Of Covalent Bonding Guided Reading and Bestseller Lists
5. Accessing Section8 2 The Nature Of Covalent Bonding Guided Reading Free and Paid eBooks
 - Section8 2 The Nature Of Covalent Bonding Guided Reading Public Domain eBooks
 - Section8 2 The Nature Of Covalent Bonding Guided Reading eBook Subscription Services
 - Section8 2 The Nature Of Covalent Bonding Guided Reading Budget-Friendly Options
6. Navigating Section8 2 The Nature Of Covalent Bonding Guided Reading eBook Formats

- ePub, PDF, MOBI, and More
 - Section8 2 The Nature Of Covalent Bonding Guided Reading Compatibility with Devices
 - Section8 2 The Nature Of Covalent Bonding Guided Reading Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section8 2 The Nature Of Covalent Bonding Guided Reading
 - Highlighting and Note-Taking Section8 2 The Nature Of Covalent Bonding Guided Reading
 - Interactive Elements Section8 2 The Nature Of Covalent Bonding Guided Reading
 8. Staying Engaged with Section8 2 The Nature Of Covalent Bonding Guided Reading
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section8 2 The Nature Of Covalent Bonding Guided Reading
 9. Balancing eBooks and Physical Books Section8 2 The Nature Of Covalent Bonding Guided Reading
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section8 2 The Nature Of Covalent Bonding Guided Reading
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Section8 2 The Nature Of Covalent Bonding Guided Reading
 - Setting Reading Goals Section8 2 The Nature Of Covalent Bonding Guided Reading
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Section8 2 The Nature Of Covalent Bonding Guided Reading
 - Fact-Checking eBook Content of Section8 2 The Nature Of Covalent Bonding Guided Reading
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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