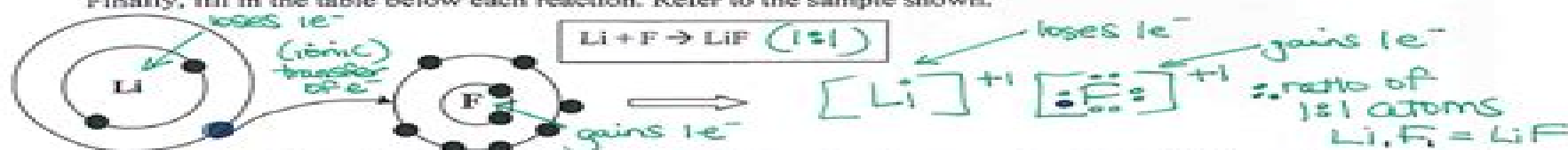


Name Solutions

Date _____ Period _____

Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers $1e^-$ to F	Li^+
F	7 ve	accepts $1e^-$ from Li	F^-

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $Li + Cl \Rightarrow LiCl$ $Li + \cdot\ddot{Cl}: \rightarrow [Li]^+ [\ddot{Cl}:]^{-1}$	Li	1 ve	donate $1e^-$	Li^+
	Cl	1 ve	accepts $1e^-$	Cl^{-1}
2) $Ca + O \Rightarrow CaO$ $Ca + \cdot\ddot{O}: \rightarrow [Ca]^{+2} [\ddot{O}:]^{-2}$	Ca	2 ve	donates $2e^-$	Ca^{+2}
	O	6 ve	accepts $2e^-$	O^{-2}
3) $Be + F \Rightarrow BeF_2$ $Be + 2 \cdot\ddot{F}: \rightarrow [Be]^{+2} [2 \cdot\ddot{F}:]^{-1}$	Be	2 ve	donates $2e^-$ to each F	Be^{+2}
	F	7 ve	each F accepts $1e^-$ each	F^{-1}
4) $Mg + S \Rightarrow MgS$ $Mg + \cdot\ddot{S}: \rightarrow [Mg]^{+2} [\ddot{S}:]^{-2}$	Mg	2 ve	donates $2e^-$ to S	Mg^{+2}
	S	6 ve	accepts $2e^-$	S^{-2}
5) $K + F \Rightarrow KF$ $K + \cdot\ddot{F}: \rightarrow [K]^+ [\ddot{F}:]^{-1}$	K	1 ve	donates $1e^-$	K^{+1}
	F	7 ve	accepts $1e^-$	F^{-1}

Section 6 Ionic Bonding Worksheet Answers

H Kauffman

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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Table of Contents Section 6 Ionic Bonding Worksheet Answers

1. Understanding the eBook Section 6 Ionic Bonding Worksheet Answers
 - The Rise of Digital Reading Section 6 Ionic Bonding Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 6 Ionic Bonding Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 6 Ionic Bonding Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 6 Ionic Bonding Worksheet Answers
 - Personalized Recommendations
 - Section 6 Ionic Bonding Worksheet Answers User Reviews and Ratings
 - Section 6 Ionic Bonding Worksheet Answers and Bestseller Lists
5. Accessing Section 6 Ionic Bonding Worksheet Answers Free and Paid eBooks
 - Section 6 Ionic Bonding Worksheet Answers Public Domain eBooks
 - Section 6 Ionic Bonding Worksheet Answers eBook Subscription Services
 - Section 6 Ionic Bonding Worksheet Answers Budget-Friendly Options

6. Navigating Section 6 Ionic Bonding Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 6 Ionic Bonding Worksheet Answers Compatibility with Devices
 - Section 6 Ionic Bonding Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 6 Ionic Bonding Worksheet Answers
 - Highlighting and Note-Taking Section 6 Ionic Bonding Worksheet Answers
 - Interactive Elements Section 6 Ionic Bonding Worksheet Answers
8. Staying Engaged with Section 6 Ionic Bonding Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 6 Ionic Bonding Worksheet Answers
9. Balancing eBooks and Physical Books Section 6 Ionic Bonding Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 6 Ionic Bonding Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section 6 Ionic Bonding Worksheet Answers
 - Setting Reading Goals Section 6 Ionic Bonding Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section 6 Ionic Bonding Worksheet Answers
 - Fact-Checking eBook Content of Section 6 Ionic Bonding Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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