

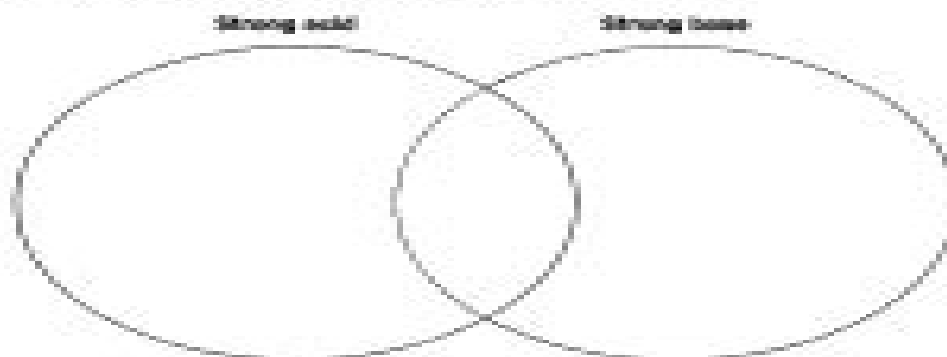
Chapter 8 Solutions, Acids, and Bases

Section 8.4 Strength of Acids and Bases**(pages 246–249)**

This section explains how to describe acids and bases in terms of both concentration and strength.

Reading Strategy (page 246)

Comparing and Contrasting As you read, complete the diagram by comparing and contrasting acids and bases. For more information on this reading strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

**The pH Scale** (page 247)

1. What is the name of the number scale chemists use to describe the concentration of hydronium ions in a solution?

2. The pH scale ranges from _____ to _____.
3. Circle the letter that indicates the pH of a neutral solution.
a. 0
b. 5
c. 7
d. 12
4. Water is neutral because it contains small but equal concentrations of _____ and _____.
5. Is the following sentence true or false? The higher the pH value of a solution, the greater the H_3O^+ ion concentration is.

6. If you add acid to pure water, the concentration of H_3O^+ _____ and the concentration of OH^- _____.

Section 8 4 Strength Of Acids And Bases

JG Myers



Section 8 4 Strength Of Acids And Bases:

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