9

CHEMICAL NAMES AND FORMULAS

SECTION 9.1 NAMING IONS (pages 253-258)

This section explains the use of the periodic table to determine the charge of an ion. It also defines polyatomic ion and gives the names and formulas for the most common polyatomic ions.

Monatomic Ions (pages 253-256)

1. What are monatomic ions?

Monatomic ions are ions consisting of only one atom.

2. How is the ionic charge of a Group 1A, 2A, or 3A ion determined?

The ionic charge is numerically equal to the group number.

3. How is the ionic charge of a Group 5A, 6A, or 7A ion determined?

The charge of an ion in Groups 5A, 6A, or 7A is determined by subtracting 8 from

the group number.

- Circle the letter of the type of element that often has more than one common ionic charge.
 - m. afficulti resettali
 - b. alkaline earth metal
 - c. transition metal
 - d. nonmetall
- The Stock system of naming transition metal cations uses a Roman numeral in parentheses to indicate the numeric value of the ionic charge.
- 6. An older naming system uses the suffix -ous to name the cation with the

charge, and the suffix -ic to name the cation with the greater charge.

7. What is a major advantage of the Stock system over the old naming system?

The Stock system gives the actual charge of the ion.

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Table of Contents Section Review Answers Chemical Names And Formulas

- 1. Understanding the eBook Section Review Answers Chemical Names And Formulas
 - The Rise of Digital Reading Section Review Answers Chemical Names And Formulas
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Section Review Answers Chemical Names And Formulas
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section Review Answers Chemical Names And Formulas
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Section Review Answers Chemical Names And Formulas
 - Personalized Recommendations
 - Section Review Answers Chemical Names And Formulas User Reviews and Ratings

- Section Review Answers Chemical Names And Formulas and Bestseller Lists
- 5. Accessing Section Review Answers Chemical Names And Formulas Free and Paid eBooks
 - Section Review Answers Chemical Names And Formulas Public Domain eBooks
 - Section Review Answers Chemical Names And Formulas eBook Subscription Services
 - Section Review Answers Chemical Names And Formulas Budget-Friendly Options
- 6. Navigating Section Review Answers Chemical Names And Formulas eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Section Review Answers Chemical Names And Formulas Compatibility with Devices
 - Section Review Answers Chemical Names And Formulas Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section Review Answers Chemical Names And Formulas
 - Highlighting and Note-Taking Section Review Answers Chemical Names And Formulas
 - Interactive Elements Section Review Answers Chemical Names And Formulas
- 8. Staying Engaged with Section Review Answers Chemical Names And Formulas
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section Review Answers Chemical Names And Formulas
- 9. Balancing eBooks and Physical Books Section Review Answers Chemical Names And Formulas
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section Review Answers Chemical Names And Formulas
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section Review Answers Chemical Names And Formulas
 - Setting Reading Goals Section Review Answers Chemical Names And Formulas
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section Review Answers Chemical Names And Formulas
 - Fact-Checking eBook Content of Section Review Answers Chemical Names And Formulas
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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