

2.1

CHEMICAL REACTIONS

Section Review

Objectives

1. Measure enthalpy using a calorimeter.
2. Identify and predict chemical reactions using phase diagrams.
3. Apply the law of conservation of energy to chemical reactions.

Vocabulary

- | | | |
|--------------------|---------------|----------------------|
| 1. endothermic | 2. exothermic | 3. enthalpy |
| 4. enthalpy change | 5. product | 6. chemical equation |

Part A Completion

Write the coefficient outside the box of each reactant or product, and write the correct phase symbol for each substance, using the phases given in parentheses. Write the balanced chemical equation.

- | | |
|---|-----------|
| 1. Aluminum reacts with oxygen to form aluminum oxide. _____ | 6. _____ |
| 2. Carbon, hydrogen, and oxygen combine to form carbon dioxide and water. _____ | 7. _____ |
| 3. Nitrogen and hydrogen combine to form ammonia. _____ | 8. _____ |
| 4. Ethanol reacts with oxygen to form carbon dioxide and water. _____ | 9. _____ |
| 5. Hydrogen peroxide decomposes to form water and oxygen. _____ | 10. _____ |
| 6. Hydrogen sulfide reacts with oxygen to form water and sulfur dioxide. _____ | 11. _____ |
| 7. Copper reacts with silver nitrate to form copper(II) nitrate and silver. _____ | 12. _____ |
| 8. Magnesium reacts with hydrochloric acid to form magnesium chloride and hydrogen gas. _____ | 13. _____ |

Part B True-False

Write true or false in the space provided. If false, write the correct statement.

1. The law of conservation of mass states that matter is neither created nor destroyed in a chemical reaction.
2. The law of conservation of energy states that energy is neither created nor destroyed in a chemical reaction.
3. The law of conservation of mass states that matter is neither created nor destroyed in a chemical reaction.
4. The law of conservation of energy states that energy is neither created nor destroyed in a chemical reaction.

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