

Acids and Bases

Reinforcement Activity

Write "A" on the line if the following substance is an acid. Write "B" on the line if it is a base. If the substance is generally not an acid or a base, write "N" for neither. Most bases are metal hydroxides (OH), metal oxides (O²⁻), or ammonium (NH₄⁺) hydroxides/oxides.

- | | | |
|---|---|--|
| 1. H ₂ O _____ | 15. Ca(OH) ₂ _____ | 29. HCl _____ |
| 2. H ₂ SO ₄ _____ | 16. CaO _____ | 30. Li ₂ O _____ |
| 3. NaOH _____ | 17. HNO ₃ _____ | 31. HBr _____ |
| 4. MgO _____ | 18. C ₆ H ₁₂ O ₆ _____ | 32. NaO _____ |
| 5. HCN _____ | 19. HNO ₂ _____ | 33. HI _____ |
| 6. HClO ₄ _____ | 20. HC ₉ O ₈ O ₅ _____ | 34. HF _____ |
| 7. HClO ₃ _____ | 21. PO ₄ ³⁻ _____ | 35. K ₂ O _____ |
| 8. HClO ₂ _____ | 22. Mg ²⁺ _____ | 36. H ₂ S _____ |
| 9. SO ₄ ²⁻ _____ | 23. H ₃ C ₃ O ₃ _____ | 37. C ₃ H ₈ _____ |
| 10. HClO _____ | 24. H ₃ BO ₃ _____ | 38. C ₃ H ₈ _____ |
| 11. NaOH _____ | 25. H ₂ SO ₃ _____ | 39. KMnO ₄ _____ |
| 12. LiOH _____ | 26. SrO _____ | 40. H ₂ CO ₃ _____ |
| 13. Ba(OH) ₂ _____ | 27. KOH _____ | 41. H ₃ PO ₄ _____ |
| 14. NH ₄ OH _____ | 28. Rb ₂ O _____ | 42. Sr(OH) ₂ _____ |

Complete and balance the acid-base neutralization reactions. Write "a" under the acid (reactant), and write "b" under the base (reactant). Write "c" under the salt, which is the new ionic compound produced. Hint: carbonate ions (CO₃²⁻) will split apart to produce CO₂ gas and to help form water.

- _____ NaOH + _____ HCl →
- _____ HC₃H₇O₃ + _____ Li₂CO₃ →
- _____ H₂SO₄ + _____ Al(OH)₃ →

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JL Elias



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