

7.1

IONS

Section Review

Part A: Completion

Use this completion section to review your understanding of the concepts and terms that are fundamental to this section.

Write each answer in the space provided, or on a separate sheet of paper.

1. Fluorine within the same group of the periodic table behaves similarly because they have the same number of valence electrons.
2. The group number of an element's periodic table tells us the number of valence electrons that element has.
3. Elements that share valence electrons are also called metals that conduct.
4. The most stable chemical structure is formed by having a positive charge and a negative charge.
5. The number of valence electrons for each element is equal to the number of protons.
6. The number of Group 1 elements which have a charge of +1.
7. The number of Group 2 elements have a charge of +2.

Part B: True-False

Write True or False next to each statement. If a statement is True, or False, then True.

- True 1. The number of valence electrons for each element is equal to the number of protons.
- True 2. The chemical properties of an element are largely determined by the number of valence electrons that element has.
- True 3. An element that has a positive charge is called a metal and a negative charge is called a nonmetal.
- True 4. An element that has a positive charge is called a metal and a negative charge is called a nonmetal.
- True 5. The number of valence electrons for each element is equal to the number of protons.
- True 6. The number of valence electrons for each element is equal to the number of protons.
- True 7. The number of valence electrons for each element is equal to the number of protons.
- True 8. The number of valence electrons for each element is equal to the number of protons.
- True 9. The number of valence electrons for each element is equal to the number of protons.
- True 10. The number of valence electrons for each element is equal to the number of protons.

Section 7 1 Ions Section Review Answers

Raffaela Di Napoli



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