

$$1. \quad \frac{1}{3} \times \frac{4}{7} = ?$$

$$a) \quad \frac{4}{7}$$

$$b) \quad \frac{4}{3}$$

$$c) \quad \frac{4}{21}$$

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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0 g $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$ 4 mol $\text{Al} \times 1 \text{ mol Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 0.132 \text{ mol Al}_2\text{O}_3$ c determine the percent yield
yield actual $\times 100 = 260.0 \text{ g} \times 100 / 86.01 \text{ g} = 302.3 \text{ g}$

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c is the limiting reagent 4 33 g of h 2 are left over 13

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 CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of co is
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