

Introduction:

The laboratory method used in the experiment is titration. Titration is a method used in measuring the amount of an analytical reagent necessary to react quantitatively with the sample. Acid-base titrations are important for counting concentrations of acids and bases. The point at which the equivalence point has been reached is called the theoretical end point. The titration is detected by some physical change, such as colour change, to indicate the end point. The indicator solution used in this experiment is phenolphthalein which is a colourless in acidic solution. When the reaction reached the endpoint, the solution becomes pale pink. One or two drops of phenolphthalein are sufficient for the titration. Addition of too much indicator will necessitate an indicator blank. The indicator should give a clear visual change in the liquid being titrated once the reaction between the standard solution and the substance is practically complete. The indicators can change colour because their ions have colours that are different from the undissociated molecule (Tan, 2010). The acid and base titrations used the Arrhenius theory. This theory stated that acid are substance which produce hydrogen ions in solution and bases are substance which produce hydroxide ion in solution. In titration, standard solutions are now usually expressed in terms of molarity, M.

Molarity, M = moles of solute/volume of solutions in litres

Objective :

One of the purposes of this experiment is to demonstrate the basic laboratory technique of titration towards the students. This experiment enables a hand-on practice done by the students themselves. Thus, the students will know how to handle the materials, apparatus and the correct way to conduct the experiment as well as being careful to avoid any experimental error. Moreover, this experiment also teach the students on the correct steps to calculate percentage of substance, molarity of solution and other calculations involve throughout the experiment.

Apparatus :

Burette with stand
Pipette
1-L volumetric flask with stopper
250 mL Erlenmeyer flasks
Retort stand with clamp

Reagents :

Sodium hydroxide pellets
Potassium hydrogen phthalate (KHP)
Phenolphthalein
Unknown vinegar

Sample Introduction Lab Report

Sebastian Brünink

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Table of Contents Sample Introduction Lab Report

1. Understanding the eBook Sample Introduction Lab Report
 - The Rise of Digital Reading Sample Introduction Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Sample Introduction Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sample Introduction Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Sample Introduction Lab Report

- Personalized Recommendations
- Sample Introduction Lab Report User Reviews and Ratings
- Sample Introduction Lab Report and Bestseller Lists
- 5. Accessing Sample Introduction Lab Report Free and Paid eBooks
 - Sample Introduction Lab Report Public Domain eBooks
 - Sample Introduction Lab Report eBook Subscription Services
 - Sample Introduction Lab Report Budget-Friendly Options
- 6. Navigating Sample Introduction Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Sample Introduction Lab Report Compatibility with Devices
 - Sample Introduction Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sample Introduction Lab Report
 - Highlighting and Note-Taking Sample Introduction Lab Report
 - Interactive Elements Sample Introduction Lab Report
- 8. Staying Engaged with Sample Introduction Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Sample Introduction Lab Report
- 9. Balancing eBooks and Physical Books Sample Introduction Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sample Introduction Lab Report
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Sample Introduction Lab Report
 - Setting Reading Goals Sample Introduction Lab Report
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Sample Introduction Lab Report

- Fact-Checking eBook Content of Sample Introduction Lab Report
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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