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web when the atoms in a bond are not the same the bonding electrons are shared and the bond is electronegativities equally  
unequally non polar polar unequally polar the degree of polarity of a bond between any two atoms is determined by  
consulting a table of

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web chemical bonds hold molecules together and create temporary connections that are essential to life types of chemical  
bonds including covalent ionic and hydrogen bonds and london dispersion forces introduction living things are made up of  
atoms but in most cases those atoms aren't just floating around individually

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web explain how a molecule that contains polar bonds can be nonpolar answer problem 6 2 2 6 2 2 which of the following  
molecules and ions contain polar bonds which of these molecules and ions have dipole moments a  $\text{ClF}_5$  b  $\text{ClO}_2$  c  $\text{TeCl}_2$  d  
 $\text{TeCl}_4$  e  $\text{PCl}_3$  f  $\text{SeF}_4$  g  $\text{PH}_2$  h  $\text{XeF}_2$  answer problem 6 2 3 6 2 3

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web the relative electronegativities of the bonded atoms is  $\text{H} < \text{C} < \text{Cl}$  and so the bond moments all point toward the  $\text{Cl}$  end of the  
molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry  
such as  $\text{BF}_3$  trigonal planar  $\text{CH}_4$  tetrahedral  $\text{PF}_5$  trigonal bipyramidal and  $\text{SF}_6$

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web explain the difference between a nonpolar covalent bond a polar covalent bond and an ionic bond answer nonpolar  
covalent electronegativity difference is less than 0.4 nonmetal nonmetal close together on the periodic table

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web apr 1 2021 updated on april 01 2021 chemical bonds may be classified as being either polar or nonpolar the difference  
is how the electrons in the bond are arranged key takeaways what is a polar bond in chemistry a polar bond is a type of  
covalent bond in which the electrons forming the bond are unequally distributed