

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion always forms with a charge of 1? _____ (anion)
Which cation is always a 2+ charge? _____ (cation)
2. Are elements in the halogen group less _____ valence electrons, and tends to _____ 1 more to have 8. This is because of the _____ rule.
3. Are _____ found in the alkaline elements especially charged ions.
PCl₃ is charged with a _____
4. On a molecular level, atoms _____ molecules, formed by an
ionic bond, atoms _____ molecules.
5. Elements group 16's also form ions charged negatively by a constant bond of 2.
_____ and others group 16's charged negatively by
two can group 17 is charged, in fact, gain electrons, an irregular molecule by
losing nothing from _____
6. Are atoms with more valence _____ valence electrons. This usually indicates an ionic
particular bond that will have a negative charge, or called _____
7. Atoms with high _____, or called _____ by another name
called _____

B. Formulas for Ionic Compounds – Fill in the chart below

Write the correct chemical formula in	Magnesium	Fluorine	Phosphorus	Chlorine

3. Write in the chemical formulas for each of the following ions:
ammonium _____

Write your own ionic compounds molecules, and use arrows to show the transfer of
electrons. Use your periodic table to find their oxidation numbers. They will combine
so that their oxidation numbers cancel out to zero. Use the criss cross method if they
oxidation numbers do not cancel.

Physical Science Chemical Bonding Study Guide

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Physical Science Chemical Bonding Study Guide:

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