

Post-Lab:

Research crystals and list other information you discovered. Use additional sheet of paper if needed but be sure to staple it to the lab sheets.

1. Where did rock candy originate?

Rock candy is one of the oldest and purest forms of candy. It was originally used by pharmacists to make medicines for many kinds of illnesses.

2. What is the name of crystals that you grew?

In the field of crystallography, these are called *monoclinic* crystals. Their shape is determined by the way the individual sugar molecules fit together, which is similar to the way the shape of a pile of oranges is determined by the shape of the individual oranges and the way they stack together.

3. What made the crystals grow?

As the water evaporates, sugar crystals form on the string or stick, and the shapes that they form reflect the shape of individual sugar crystals.

4. What was the purpose of soaking the skewer or string?

The string will provide the surface on which the crystals will grow. As water evaporates from the string, small crystals of sugar will encrust the string. These tiny seed crystals provide starting points for larger crystals. Future growth will be concentrated around these points.

5. What is a super saturated solution?

When you mixed the water and sugar you made a Super Saturated Solution which means that the water could only hold the sugar if both were very hot. As the water cooled, the sugar came out of the solution and went back into sugar crystals on the skewer.

Rock Candy Lab Chemistry Answers

Michael Horton



Rock Candy Lab Chemistry Answers:

The Handy Chemistry Answer Book Justin P. Lomont, Ian C. Stewart, 2013-10-01 Don't be mixed up about chemistry! Simplify the complex chemical reactions that take place everywhere in our lives with this engaging easy to follow question and answer guide. Where would we be without atoms and compounds? Gas, liquids, solids, and plasma. Acids and bases. Bonds and reactions. Matter and energy. The Handy Chemistry Answer Book covers the building blocks of life and the universe. The secret life of atoms, how polar bears aren't actually white, why oil and water don't mix, and much, much more are revealed and explained. This informative guide covers the basics of chemistry, history, atomic structures, chemical bonds, and reactions, organic and inorganic chemistry, to more advanced material, nuclear chemistry, biochemistry, physical and theoretical chemistry, by answering nearly 1,000 common chemistry questions including: What causes lightning? How does photosynthesis work? What are hard and soft Lewis acids and bases? What makes a fabric waterproof? What are the twelve principles of green chemistry? When did alchemists finally abandon trying to make gold? What is Le Chatelier's principle? What do the different octane ratings mean at the gas pump? What is genetic engineering? Why is calcium important for strong bones? What is the 18 electron rule? Why does chocolate turn white as it ages? Chemical reactions that rule the world, their properties, structure, composition, behavior, and history are tackled and explained in plain English in *The Handy Chemistry Answer Book*. With many photos, illustrations, a few formulas, molecular diagrams, and other graphics, this fun fact filled tome is richly illustrated. A history of chemistry, timeline, appendices on Nobel Prize winners, a bibliography, further reading section, glossary of terms, a table of physical constants, a table of conversion factors, and extensive index add to its usefulness. Applied Chemistry Experiment Sheets Martin Mendel, 1926 *Maryland School Bulletin* Maryland. State Dept. of Education, 1921

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