

Stoichiometry Worksheet #1

1. Aluminum chloride, AlCl_3 , is used as a catalyst in various industrial reactions. It is prepared from hydrogen chloride gas and aluminum metal shavings.



1A. How many moles of AlCl_3 can be prepared from 3.5 moles of hydrogen chloride gas with an excess of aluminum?

$$3.5 \text{ mol HCl} \times \frac{2 \text{ mol AlCl}_3}{6 \text{ mol HCl}} = 1.2 \text{ mol AlCl}_3$$

Ans: 1.2 mol AlCl_3

1B. How many moles of hydrogen gas would be produced from the use of 8.5 moles of aluminum with an excess of hydrogen chloride?

Ans: 13 mol H_2

$$8.5 \text{ mol Al} \times \frac{3 \text{ mol H}_2}{2 \text{ mol Al}} = 13 \text{ mol H}_2$$

2. When dinitrogen pentoxide, N_2O_5 , a white solid, is heated, it decomposes to nitrogen dioxide and oxygen.



2A. How many moles of nitrogen dioxide can be formed from the decomposition of 1.25 g of N_2O_5 ?

$$1.25 \text{ g N}_2\text{O}_5 \times \frac{1 \text{ mol N}_2\text{O}_5}{108.02 \text{ g N}_2\text{O}_5} \times \frac{4 \text{ mol NO}_2}{2 \text{ mol N}_2\text{O}_5} = 0.0231 \text{ mol NO}_2$$

Ans: 0.0231 mol NO_2

2B. How many grams of oxygen can be formed from the decomposition of 2.3 g of N_2O_5 ?

$$2.3 \text{ g N}_2\text{O}_5 \times \frac{1 \text{ mol N}_2\text{O}_5}{108.02 \text{ g N}_2\text{O}_5} \times \frac{1 \text{ mol O}_2}{2 \text{ mol N}_2\text{O}_5} \times \frac{32.00 \text{ g O}_2}{1 \text{ mol O}_2} = 0.34 \text{ g O}_2$$

Ans: 0.34 g O_2

3. Chlorine is prepared from sodium chloride by electrochemical decomposition. Formerly chlorine was produced by heating hydrochloric acid with pyrolusite [manganese dioxide or manganese(IV) oxide, MnO_2], a common manganese ore. Small amounts of chlorine may be prepared in the laboratory by the same reaction.



3A. How many grams of HCl react with 5.00 g of MnO_2 , according to the equation?

$$5.00 \text{ g MnO}_2 \times \frac{1 \text{ mol MnO}_2}{86.94 \text{ g MnO}_2} \times \frac{4 \text{ mol HCl}}{1 \text{ mol MnO}_2} \times \frac{36.46 \text{ g HCl}}{1 \text{ mol HCl}} = 8.39 \text{ g HCl}$$

Ans: 8.39 g HCl

3B. If a chemist wanted to prepare 100. g of chlorine, how many grams of MnO_2 are needed, assuming there is more than enough hydrochloric acid?

Ans: 123 g MnO_2

$$100. \text{ g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.90 \text{ g Cl}_2} \times \frac{1 \text{ mol MnO}_2}{1 \text{ mol Cl}_2} \times \frac{86.94 \text{ g MnO}_2}{1 \text{ mol MnO}_2} = 123 \text{ g MnO}_2$$

3C. How many molecules of water are produced from the reaction of 5.0 g of HCl?

Ans: 4.1×10^{23} molec. H_2O

$$5.0 \text{ g HCl} \times \frac{1 \text{ mol HCl}}{36.46 \text{ g HCl}} \times \frac{2 \text{ mol H}_2\text{O}}{4 \text{ mol HCl}} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 4.1 \times 10^{23} \text{ molecules H}_2\text{O}$$

4. Sodium is a soft, reactive metal that instantly reacts with water to give hydrogen gas and a solution of sodium hydroxide,



4A. How many grams of sodium metal are needed to give 7.81 g of hydrogen by this reaction?

Ans: 178 g Na

$$7.81 \text{ g H}_2 \times \frac{1 \text{ mol H}_2}{2.02 \text{ g H}_2} \times \frac{2 \text{ mol Na}}{1 \text{ mol H}_2} \times \frac{22.99 \text{ g Na}}{1 \text{ mol Na}} = 178 \text{ g Na}$$

4B. How many sodium atoms are needed to react with 1.25×10^{24} molecules of water?

Ans: 1.25×10^{24} atoms Na

$$1.25 \times 10^{24} \text{ molecules H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{6.02 \times 10^{23} \text{ molecules H}_2\text{O}} \times \frac{2 \text{ mol Na}}{2 \text{ mol H}_2\text{O}} \times \frac{6.02 \times 10^{23} \text{ atoms Na}}{1 \text{ mol Na}} =$$

Power Notes Introduction To Stoichiometry Worksheet

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Schneider**



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James Patrick Abulencia, Louis Theodore, 2015-03-27 This is a unique book with nearly 1000 problems and 50 case studies on open ended problems in every key topic in chemical engineering that helps to better prepare chemical engineers for the future The term open ended

problem basically describes an approach to the solution of a problem and or situation for which there is not a unique solution The Introduction to the general subject of open ended problems is followed by 22 chapters each of which addresses a traditional chemical engineering or chemical engineering related topic Each of these chapters contain a brief overview of the subject matter of concern e g thermodynamics which is followed by sample open ended problems that have been solved by the authors employing one of the many possible approaches to the solutions This is then followed by approximately 40 45 open ended problems with no solutions although many of the authors solutions are available for those who adopt the book for classroom or training purposes A reference section is included with the chapter s contents Term projects comprised of 12 additional chapter topics complement the presentation This book provides academic industrial and research personnel with the material that covers the principles and applications of open ended chemical engineering problems in a thorough and clear manner Upon completion of the text the reader should have acquired not only a working knowledge of the principles of chemical engineering but also and more importantly experience in solving open ended problems What many educators have learned is that the applications and implications of open ended problems are not only changing professions but also are moving so fast that many have not yet grasped their tremendous impact The book drives home that the open ended approach will revolutionize the way chemical engineers will need to operate in the future

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Introduction to Material and Energy Balances Gintaras V. Reklaitis, Daniel R. Schneider, 1983 A thorough introduction to balance equation concepts Geared for the course offered to chemical engineering majors in their sophomore year Develops a framework for the analysis of flowsheet problem information with extensive use of degree of freedom analysis Presents systematic approaches for manual and computer aided solution of full scale balance problems Provides a detailed development of the structure properties and interrelationships of species and element balances based on the algebraic view of reaction stoichiometry and the rate of reaction concept

Unit Operations in Environmental Engineering Louis Theodore, R. Ryan Dupont, Kumar Ganesan, 2017-08-29 The book presents the principles of unit operations as well as the application of these principles to real world problems The authors have written a practical introductory text exploring the theory and applications of unit operations for environmental engineers that is a comprehensive update to Linvil Rich s 1961 classic work Unit Operations in Sanitary Engineering The book is designed to serve as a training tool for those individuals pursuing degrees that include courses on unit operations Although the literature is inundated with publications in this area emphasizing theory and theoretical derivations the goal of this book is to present the subject from a strictly pragmatic introductory point of view particularly for those individuals involved with environmental engineering This book is concerned with unit operations fluid flow heat transfer and mass transfer Unit operations by definition are physical processes although there are some that include chemical and biological reactions The unit operations approach allows both the practicing engineer and student to compartmentalize the various operations that constitute a process and emphasizes introductory engineering principles so that the reader can then satisfactorily predict the performance of the various unit operations equipment This is a definitive work on Unit Operations one of the most important subjects in environmental engineering today It is an excellent reference well written easily read and comprehensive I believe the book will serve well those working in engineering disciplines including those beyond just environmental and chemical engineering Bottom line A must for any technical library Kenneth J Skipka CCM

Resources in Education, 1995 *Stoichiometry Unit Project* Luann Marie Decker, 1998

Chemistry for the IB Diploma Geoff Neuss, 2001 This concise guide provides the content needed for the Chemistry IB diploma at both Standard and Higher Level It follows the structure of the IB Programme exactly and includes all the options Each topic is presented on its own page for clarity Higher Level material is clearly indicated and there are plenty of practice questions The text is written with an awareness that English might not be the reader s first language

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