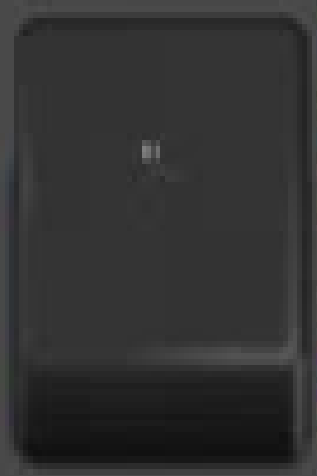
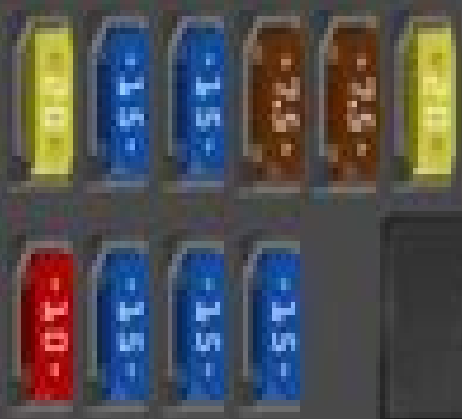
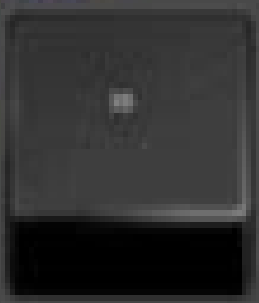


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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 230 \text{ kJ}$
 $2 \text{ H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g}) \rightarrow \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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web thus for reaction 9.2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta h = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta h_{\text{products}}$

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions

1 $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{aq} + \text{OH}^- \text{aq}$ $\Delta H = -44.5 \text{ kJ/mol}$

2 $\text{NaOH (s)} + \text{H}^+ \text{aq} \rightarrow \text{Na}^+ \text{aq} + \text{H}_2\text{O (l)}$ $\Delta H = -57.3 \text{ kJ/mol}$

3 $\text{NaOH (s)} + \text{HCl (aq)} \rightarrow \text{Na}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{H}_2\text{O (l)}$ $\Delta H = -57.3 \text{ kJ/mol}$