

7-2 Review and Reinforcement

Covalent Bonding

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line provided.

- _____ F, covalent 1. A group of atoms united by ionic bonds is called a molecule.
- _____ T 2. A covalent bond is formed by a shared pair of electrons
- _____ F, four 3. A double covalent bond consists of two shared electrons.
- _____ T 4. A molecular formula tells you how many atoms are in a single molecule of the compound.
- _____ F, MF 5. The empirical formula for a molecule specifies which atoms are bonded to each other in the molecule.
- _____ F, lone pair 6. A pair of electrons not involved in bonding is called a shared electron pair.
- _____ F, equally 7. In nonpolar covalent bonds, the electrons are shared unequally between two atoms.
- _____ F, Triple 8. Lewis structures use a triple dash to represent a double bond.

Use the information in Figure 7-19 of the textbook to calculate the electronegativity differences between each pair of elements listed in the chart below. Then indicate whether the bond would be nonpolar covalent, polar covalent, or ionic. Write your answers in the chart.

		Electronegativity Differences	Type of Bond
9.	H-O		
10.	C-H		
11.	K-F		
12.	N-H		
13.	Na-F		
14.	O-Cl		

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Table of Contents Science 7 2 Review And Reinforce Answers

1. Understanding the eBook Science 7 2 Review And Reinforce Answers
 - The Rise of Digital Reading Science 7 2 Review And Reinforce Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Science 7 2 Review And Reinforce Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Science 7 2 Review And Reinforce Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Science 7 2 Review And Reinforce Answers
 - Personalized Recommendations
 - Science 7 2 Review And Reinforce Answers User Reviews and Ratings
 - Science 7 2 Review And Reinforce Answers and Bestseller Lists
5. Accessing Science 7 2 Review And Reinforce Answers Free and Paid eBooks
 - Science 7 2 Review And Reinforce Answers Public Domain eBooks
 - Science 7 2 Review And Reinforce Answers eBook Subscription Services
 - Science 7 2 Review And Reinforce Answers Budget-Friendly Options
6. Navigating Science 7 2 Review And Reinforce Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Science 7 2 Review And Reinforce Answers Compatibility with Devices
 - Science 7 2 Review And Reinforce Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Science 7 2 Review And Reinforce Answers
 - Highlighting and Note-Taking Science 7 2 Review And Reinforce Answers
 - Interactive Elements Science 7 2 Review And Reinforce Answers
 8. Staying Engaged with Science 7 2 Review And Reinforce Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Science 7 2 Review And Reinforce Answers
 9. Balancing eBooks and Physical Books Science 7 2 Review And Reinforce Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Science 7 2 Review And Reinforce Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Science 7 2 Review And Reinforce Answers
 - Setting Reading Goals Science 7 2 Review And Reinforce Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Science 7 2 Review And Reinforce Answers
 - Fact-Checking eBook Content of Science 7 2 Review And Reinforce Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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