Science 30 UNIT D: Energy and the Environment

Harvesting Chemical Energy and Hess's Law Assignment

Please use the "Heat of Formation" values found in the Science 30 data booklet. Remember that there are different values for water vapour $[H_2O(g)]$ and liquid water $[H_2O(l)]$.

- The following balanced chemical equations are for based on the combustions of various fuels.
 Calculate the heats of combustion for each reaction
 - a. C₂H₅OH (I) + 3 O₂ (g) → 2 CO₂ (g) + 3 H₂O (g)

b. $2 C_4H_{10}(g) + 13 O_2(g) \rightarrow 8 CO_2(g) + 10 H_2O(g)$

c. $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(g)$

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