

Stoichiometry Review: Test Review



1. What is stoichiometry? [] Study of quantitative relationships in chemical reaction.
2. On which law is stoichiometry based? [] Law of conservation of mass.
3. According to the law of conservation of mass, the mass of the [] ^{reactants} must [equal] the mass of the ^{products}.
4. What is a mole ratio? [] fractions/ratio comparing two substances.
5. Given the following balanced equation: $2\text{Al} + 3\text{Br}_2 \rightarrow 2\text{AlBr}_3$, determine the mole ratio used for calculating moles of aluminum bromide from moles of bromine. [] $2\text{mol AlBr}_3 / 3\text{mol Br}_2$
6. What must you have before beginning any stoichiometric problem? [] balanced equation
7. What is a limiting reactant? [] runs out
8. What is an excess reactant? [] left over
9. A booklet is prepared from 2 covers, 3 staples, and 20 pages. Assume we have 60 covers, 120 staples, and 400 pages.
 - a. What is the largest number of complete booklets that can be made? [20]
 - b. Which is the "limiting reactant" in this system? [pages]
10. How many moles of lithium chloride will be formed by the reaction of chlorine with 0.046 mol of lithium bromide in the following reaction?
 $[2] \text{LiBr} + [1] \text{Cl}_2 \rightarrow [2] \text{LiCl} + [1] \text{Br}_2$

$$\frac{0.046 \text{ mol LiBr}}{2 \text{ mol LiBr}} \times \frac{2 \text{ mol LiCl}}{2 \text{ mol LiBr}} = 0.046 \text{ mol LiCl}$$

11. Hydrogen peroxide breaks down, releasing oxygen, in the following reaction. What mass of oxygen is produced when 1.840 mol of H_2O_2 decomposes?

$$\frac{1.840 \text{ mol H}_2\text{O}_2}{2 \text{ mol H}_2\text{O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}_2} \times \frac{32.09 \text{ g O}_2}{1 \text{ mol O}_2} = 29.44 \text{ g O}_2$$

12. Hydrogen is generated by passing hot steam over iron, which oxidizes to form Fe_2O_3 , in the following equation. If 625 g of Fe_2O_3 is produced in the reaction, how many moles of hydrogen are produced at the same time?

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Y Pai



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