

Lesson 2. Nuclear Chemistry & Energy

Outline

- 2.1 Nuclear Reactions and Radioactivity
- 2.2 Nuclear Stability
- 2.3 Kinetics of Radioactive Decay
- 2.4 Energetics of Nuclear Reactions
- 2.5 Fusion, Transmutation, Fission
- 2.6 Interaction of Radiation and Matter

Nuclear Chemistry

Nuclear chemistry is a study of nuclear reactions. Nuclear reactions are changes in matter originating in the nucleus of an atom. When a nuclei change spontaneously, it emits radiation and that said is for radioactivity. Radioactive elements are widely used in medicine: diagnostic tools and as a means of treatment especially for cancer. They are also used as test substances to determine the mechanisms of chemical reactions, to trace the movement of atoms in biological systems, and as date important historical artifacts. Nuclear reactions are also used both to generate electricity and to create weapons of massive destruction.

This lesson discussed nuclear reactions and its applications.

2.1 Nuclear Reactions and Radioactivity

Nuclear Reactions

One of the nuclear reactions in the universe is the formation of carbon-14, which is a result of collision of gas molecules in the atmosphere with cosmic rays. Cosmic rays are particles travelling at high speeds. Majority of cosmic rays are atomic nuclei. Roughly, 80% of cosmic rays are hydrogen nuclei and about 10% are helium nuclei. Some cosmic rays emanate from the sun and others originate outside the solar system. The energies of cosmic rays are far higher than chemical energies. (Chemical energies are usually expressed in kilojoules per mole but that of cosmic rays are expressed in electron-volts). One eV is equal to 1.6×10^{-19} Joule.

As cosmic rays enter the atmosphere, it collided with gas molecules. These collisions initiate nuclear reactions in the atmosphere. One of this nuclear reactions leads to the formation of the radioactive carbon-14. Carbon-14 is unstable isotope of carbon. It undergoes spontaneous radioactive decay or disintegration, spitting particles from the nucleus. Carbon-14 is formed when a free neutron is absorbed by nitrogen nucleus resulting to carbon nucleus and proton. Nitrogen nucleus and neutron are the reactants and carbon nucleus and proton are the products. The nuclear reaction is summarized



Note that the reactants and products of a nuclear reaction are summarised as follows: atomic notation, ${}^A_Z\text{X}$, where X is the symbol of the atom, A is the mass number and Z is the atomic number or nuclear charge. Similar notation is used for subatomic particles, neutrons (${}^1_0\text{n}$), protons (${}^1_1\text{p}$), and electrons (${}^0_{-1}\text{e}$).

Nuclear reactions also obey conservation rules, hence, nuclear reactions are balanced. Balancing nuclear equation is with respect to both mass number and nuclear charge. For the nuclear equation to form carbon-14, the sum of the mass numbers for each side of the equation is 14, and the sum of the nuclear charges on each side of the equation is 7, therefore the nuclear equation is balanced.

Review Section 2 Nuclear Chemistry

Christian G. Meyer



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