

CHAPTER 17

RADIOACTIVITY AND NUCLEAR CHEMISTRY

Practice Problems

1. (a) Write a nuclear equation for alpha decay of Po-210 .



- (b) Write a nuclear equation for the beta decay of Th-231 .



- (c) Write a nuclear equation for positron emission by Co-57 .



2. A 10.0-g sample of a radioactive nuclide is administered to a patient for medical diagnosis of liver function. If the nuclide has a half-life of 5.3 hours, how much of the nuclide remains in the patient after 4.0 days?

$$\text{4 days} = \frac{4 \text{ days}}{24 \text{ hr/day}} = \frac{96 \text{ hr}}{24 \text{ hr/day}} = 4 \text{ days} \quad \frac{1 \text{ day}}{24 \text{ hr/day}} = \frac{1 \text{ day}}{24 \text{ hr/day}} = 0.0417 \text{ day/hr}$$

- (a) $\frac{1 \text{ day}}{24 \text{ hr/day}} = \frac{1 \text{ day}}{24 \text{ hr/day}} = 0.0417 \text{ day/hr}$ (b) $\frac{1 \text{ day}}{24 \text{ hr/day}} = \frac{1 \text{ day}}{24 \text{ hr/day}} = 0.0417 \text{ day/hr}$ (c) $\frac{1 \text{ day}}{24 \text{ hr/day}} = \frac{1 \text{ day}}{24 \text{ hr/day}} = 0.0417 \text{ day/hr}$ (d) $\frac{1 \text{ day}}{24 \text{ hr/day}} = \frac{1 \text{ day}}{24 \text{ hr/day}} = 0.0417 \text{ day/hr}$ (e) $\frac{1 \text{ day}}{24 \text{ hr/day}} = \frac{1 \text{ day}}{24 \text{ hr/day}} = 0.0417 \text{ day/hr}$

3. A mammoth skeleton has a carbon-14 content of 18.3% of that found in living organisms. When did the mammoth live?

The age of the mammoth is equal to the half-life of C-14 , which is approximately

$$5730 \text{ yrs}$$

Review Nuclear Chemistry Answer Key Section 22

CL Gary



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