

Honors Chem - Stoichiometry Quiz Summary Review

Stoichiometry

- the study of quantitative relationships in a chemical equation
- used to predict how much product will be formed
- used to determine the amount of reactants needed

EXAMPLE



The coefficients (underlined) of a balanced equation indicate the mole ratio of substances in a reaction.

How many moles of CO_2 will be produced when 8 moles of C_3H_8 combusts completely?

Step 1: Plan

We have to go from moles $\text{CO}_2 \rightarrow$ moles C_3H_8

Step 2: Calculate

$$\begin{aligned} 8 \text{ mol C}_3\text{H}_8 &\times 3 \text{ mol CO}_2 / 1 \text{ mol C}_3\text{H}_8 \\ &= 24 \text{ mol CO}_2 \end{aligned}$$

Limiting Reactant

- completely used by the reaction
- no remaining reactant at the end of the reaction
- determines the maximum amount of product formed
- balloon experiment flask D: vinegar (not enough to make the balloon any bigger)

Excess Reactant

- reactant left over at the end of the reaction
- does not completely react
- balloon experiment flask D: baking soda (was left over in the flask)

EXAMPLE

13.8 g of baking soda (NaHCO_3) was added to 0.0833 moles of vinegar ($\text{HC}_2\text{H}_3\text{O}_2$) according to the balanced equation below:



What is the maximum volume of CO_2 (g) that could be produced at STP?

(Useful GMFS (g/mol): NaHCO_3 - 84.0; $\text{HC}_2\text{H}_3\text{O}_2$ - 60.0)

Step 1: Plan

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SA Adler



Review Chemistry Honors Unit Stoichiometry:

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Applied Science & Technology Index, 1966 Stanford University Bulletin Stanford University, 1999 *Catalogs of Courses* University of California, Berkeley, 1990 Includes general and summer catalogs issued between 1878 1879 and 1995 1997 *General Catalog* Iowa State University, 1989 **Schedule of Classes** University of California, San Diego, 1996 *Current Index to Journals in Education*, 1989

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Formulation and Stoichiometry Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the

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