

# Acids and Bases

## Reinforcement Activity

Write "A" on the line if the following substance is an acid. Write "B" on the line if it is a base. If the substance is generally not an acid or a base, write "N" for neither. Most bases are metal hydroxides (OH), metal oxides (O<sup>2-</sup>), or ammonium (NH<sub>4</sub><sup>+</sup>) hydroxides/oxides.

- |                                         |                                                         |                                          |
|-----------------------------------------|---------------------------------------------------------|------------------------------------------|
| 1. H <sub>2</sub> O _____               | 15. Ca(OH) <sub>2</sub> _____                           | 29. HCl _____                            |
| 2. H <sub>2</sub> SO <sub>4</sub> _____ | 16. CaO _____                                           | 30. Li <sub>2</sub> O _____              |
| 3. NaOH _____                           | 17. HNO <sub>3</sub> _____                              | 31. HBr _____                            |
| 4. MgO _____                            | 18. C <sub>6</sub> H <sub>12</sub> O <sub>6</sub> _____ | 32. NaO _____                            |
| 5. HCN _____                            | 19. HNO <sub>2</sub> _____                              | 33. HI _____                             |
| 6. HClO <sub>4</sub> _____              | 20. HC <sub>9</sub> O <sub>8</sub> O <sub>5</sub> _____ | 34. HF _____                             |
| 7. HClO <sub>3</sub> _____              | 21. PO <sub>4</sub> <sup>3-</sup> _____                 | 35. K <sub>2</sub> O _____               |
| 8. HClO <sub>2</sub> _____              | 22. Mg <sup>2+</sup> _____                              | 36. H <sub>2</sub> S _____               |
| 9. SO <sub>4</sub> <sup>2-</sup> _____  | 23. H <sub>3</sub> C <sub>7</sub> O <sub>8</sub> _____  | 37. C <sub>3</sub> H <sub>8</sub> _____  |
| 10. HClO _____                          | 24. H <sub>3</sub> BO <sub>3</sub> _____                | 38. C <sub>3</sub> H <sub>8</sub> _____  |
| 11. NaOH _____                          | 25. H <sub>2</sub> SO <sub>3</sub> _____                | 39. KMnO <sub>4</sub> _____              |
| 12. LiOH _____                          | 26. SrO _____                                           | 40. H <sub>2</sub> CO <sub>3</sub> _____ |
| 13. Ba(OH) <sub>2</sub> _____           | 27. KOH _____                                           | 41. H <sub>3</sub> PO <sub>3</sub> _____ |
| 14. NH <sub>4</sub> OH _____            | 28. Rb <sub>2</sub> O _____                             | 42. Sr(OH) <sub>2</sub> _____            |

Complete and balance the acid-base neutralization reactions. Write "a" under the acid (reactant), and write "b" under the base (reactant). Write "s" under the salt, which is the new ionic compound produced. Hint: carbonate ions (CO<sub>3</sub><sup>2-</sup>) will split apart to produce CO<sub>2</sub> gas and to help form water.

- \_\_\_\_\_ NaOH + \_\_\_\_\_ HCl →
- \_\_\_\_\_ HC<sub>3</sub>H<sub>7</sub>O<sub>3</sub> + \_\_\_\_\_ Li<sub>2</sub>CO<sub>3</sub> →
- \_\_\_\_\_ H<sub>2</sub>SO<sub>4</sub> + \_\_\_\_\_ Al(OH)<sub>3</sub> →

# Review And Reinforcement Acids And Bases Answers

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