

CHAPTER 14 REVIEW*Acids and Bases***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

- 1.**
- Name the following compounds as acids:

_____ sulfuric acid _____	a. H_2SO_4
_____ sulfurous acid _____	b. H_2SO_3
_____ hydrosulfuric acid _____	c. H_2S
_____ perchloric acid _____	d. HClO_4
_____ hydrocyanic acid _____	e. hydrogen cyanide

- 2.**
- _____
- H_2S
- Which (if any) of the acids mentioned in item
- 1**
- are binary acids?

- 3.**
- Write formulas for the following acids:

_____ HNO_2 _____	a. nitrous acid
_____ HBr _____	b. hydrobromic acid
_____ H_3PO_4 _____	c. phosphoric acid
_____ CH_3COOH _____	d. acetic acid
_____ HClO _____	e. hypochlorous acid

- 4.**
- Calcium selenate has the formula
- CaSeO_4
- .

_____ H_2SeO_4 _____	a. What is the formula for selenic acid?
_____ H_2SeO_3 _____	b. What is the formula for selenous acid?

- 5.**
- Use an activity series to identify two metals that will not generate hydrogen gas when treated with an acid.

Choose from Cu, Ag, Au, Pt, Pd, or Hg.

- 6.**
- Write balanced chemical equations for the following reactions of acids and bases:

- a. aluminum metal with dilute nitric acid



- b. calcium hydroxide solution with acetic acid



Review Acids And Bases Holt Chemistry

Peter John Turnbull Morris



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1. Understanding the eBook Review Acids And Bases Holt Chemistry
 - The Rise of Digital Reading Review Acids And Bases Holt Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Acids And Bases Holt Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Acids And Bases Holt Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Acids And Bases Holt Chemistry
 - Personalized Recommendations
 - Review Acids And Bases Holt Chemistry User Reviews and Ratings
 - Review Acids And Bases Holt Chemistry and Bestseller Lists
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 - Review Acids And Bases Holt Chemistry Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Review Acids And Bases Holt Chemistry
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- 8. Staying Engaged with Review Acids And Bases Holt Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Acids And Bases Holt Chemistry
- 9. Balancing eBooks and Physical Books Review Acids And Bases Holt Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Acids And Bases Holt Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review Acids And Bases Holt Chemistry
 - Setting Reading Goals Review Acids And Bases Holt Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Acids And Bases Holt Chemistry
 - Fact-Checking eBook Content of Review Acids And Bases Holt Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
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