## Acids & Bases: Assessment Review

Adjusted on

Skill - The student will be able to identify the acid, base, conjugate acid, and conjugate base in an acid - hase reaction.

Write a chemical equation showing how HCO, behaves as an acid when dissolved in water.
 Label the acid, base, conjugate acid & conjugate base.

 Write a chemical equation showing how ammonia IMH J behaves as a base when dissolved in water. Label the ecid, base, conjugate ecid & conjugate base.

Sibill - The student will be able to calculate with the (M+) or (GM-) from the ion product 50w) or pM from pGM (a one-step pM block calculation)

Recreate the calculation flowchast on the diagram provided: (Se careful, Its hean flipped).





- 4. What is [OH] if a solution has [H] = 7.1 x 10" M? Is the solution acidic, basic, or neutral?
- 5. Calculate D4°1 in a solution of 50m1 = 5.2 x 3.0 MM. In the solution acidic, basic, or neutral?
- Calculate the [H\*] in a solution that shows a pH of 11.70, listhe solution acids, basic, or neutral?
- 7. What is the prim of a solution of COH1 = 5.0 x 10 ° M7 is the solution acidic, basic, or neutral?

# **Review Acids And Bases Answer Key Holt**

**Holt, Rinehart and Winston Staff** 

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