

Chapter 10 Nuclear Chemistry

7. How is the product isotope different from the reactant isotope in alpha decay? The product isotope has two fewer protons and two fewer neutrons than the reactant isotope.
8. Circle the letters that identify each event that takes place during beta decay.
- ☐ A proton decomposes into a neutron and an electron.
 - ☒ A neutron decomposes into a proton and an electron.
 - ☒ An electron is emitted from the nucleus.
 - ☐ A neutron is emitted from the nucleus.
9. Why are beta particles more penetrating than alpha particles? Beta particles have a smaller mass and a faster speed.
10. Is the following sentence true or false? All nuclear radiation consists of charged particles. false
11. What is a gamma ray? A gamma ray is a penetrating ray of energy emitted by an unstable nucleus.
12. How fast do gamma rays travel through space? Gamma rays travel through space at the speed of light.
13. Complete the following table about nuclear radiation.

Characteristics of Nuclear Radiation			
Radiation Type	Charge	Mass (amu)	Usually Stopped By
Alpha particle	- 1	4	Paper or clothing
Beta particle	1 -	1 1836	Aluminum sheet
Gamma ray	0	0	Several meters of concrete

Effects of Nuclear Radiation (pages 296–297)

14. How does nuclear radiation affect atoms? Nuclear radiation can ionize atoms.
15. Is the following sentence true or false? One potential danger of radon gas is that prolonged exposure to it can lead to lung cancer. true

Detecting Nuclear Radiation (page 297)

16. Name two devices that are used to detect nuclear radiation.
- Geiger counters
 - Film badges

Review Nuclear Chemistry Section 1 Answer Key

Jin-Ying Zhang



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hscn present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species
 $\text{mol Fe}^{3+} = 1.70 \times 10^{-5} \text{ mol}$
 $\text{mol Fe}^{3+} = 8.50 \times 10^{-4} \text{ mol}$

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