

Chem 1B Arasasingham
Midterm 2 Review Key

Datta

1. Calculate the temperature of an ideal gas that has an average kinetic energy of 6.73 kJ/mol.

Equation to use: $KE_{avg} = \frac{3}{2} (RT)$

$R = 8.314 \text{ J/K}\cdot\text{mol}$

$KE_{avg} = 6.73 \text{ kJ/mol} = 6730 \text{ J/mol}$

$T = ?$

$T = \frac{2}{3} \left(\frac{KE_{avg}}{R} \right) = \frac{2}{3} \left(\frac{6730 \text{ J/mol}}{8.314 \text{ J/K}\cdot\text{mol}} \right) = 539 \text{ K}$

2. Find the root mean square speed of gaseous nitrogen molecules at 95 degrees Celsius.

Given:

$U_{rms} = ?$

$T = 95 \text{ C} = 273 + 95 = 368 \text{ K}$

$R = 8.314 \text{ J/K}\cdot\text{mol}$

$M = \text{molar mass of nitrogen gas (N}_2\text{)} = (14.007) \cdot 2 = 28.014 \text{ g/mol N}_2$

Equation to use: $U_{rms} = \sqrt{\square}$

$U_{rms} = \sqrt{\square} = \sqrt{\square}$; remember that $1 \text{ J} = 1 \text{ kg(m/s)}^2$

Then, $\sqrt{\square} \cdot (1000 \text{ g} / 1 \text{ kg}) = \sqrt{\square} = \sqrt{\square}$

$U_{rms} = 572.4 \text{ m/s}$

3. Calculate the average translational kinetic energy ($\bar{J}$) of a At molecule that has a molecular root mean square speed of 437.2 m/s.

***note: $1 \text{ J} = 1 \text{ kg(m}^2\text{/s}^2\text{)}$

STEP 1: Find the mass of the atom (in kg, to ease calculations later).

$(1 \text{ At atom}) \left(\frac{1 \text{ mol At}}{6.022 \times 10^{23} \text{ atoms}} \right) \times \frac{209.99 \text{ g}}{1 \text{ mol At}} \times \frac{1 \text{ kg}}{10^3 \text{ g}} = 3.48 \times 10^{-25} \text{ kg}$

STEP 2: Calculate KE using the equation $KE = \frac{1}{2}(m_0)(u^2)$

$KE = \frac{1}{2}(3.48 \times 10^{-25} \text{ kg At}) \cdot (437.2 \text{ m/s})^2 = 3.33 \times 10^{-18} \text{ J}$

4. What is the molar mass (g/mol) of a gas whose root mean square speed is 621 m/s at 436 K?

$U_{rms} = \sqrt{\square}$

$T = 436 \text{ K}$

$R = 8.314 \text{ J/K}\cdot\text{mol}$

$U_{rms} = 621 \text{ m/s}$

$M = ? \text{ (solve)}$

$M = (3RT) / (U_{rms})^2 = [(3) \cdot (8.314 \text{ kgm}^2\text{/s}^2\text{/K}\cdot\text{mol}) \cdot (436 \text{ K})] / (621 \text{ m/s})^2 = 0.0282 \text{ kg/mol}$

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Katrin Zwirgmaier



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