



NATIONAL
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

MARKS: 100
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use permanent marker for writing and drawing.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fus} for 1-hexanol in kJ mol⁻¹ and in $^{\circ}\text{C}$ respectively.

Compound	Molecular mass (M _r)	Boiling point (T _b)
A	102	163
B	102	17

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B? or C? and list one of the reasons. (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and its boiling point of these molecules. (2)
- 3.1.4 What kind of the compound A or B is 1-hexanol? Explain how you arrived at your answer. (2)
- 3.1.5 Which compound (A or B) will have the higher vapour pressure? (1)
- 3.1.6 Give a reason for your answer. (1)
- 3.1.7 Which compound (A or B) will have the higher surface tension? (1)
- 3.1.8 Give a reason for the answer to 3.1.7 relating to the space the table. (1)

Compound C (M_r 102) will show in the table, has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 163 $^{\circ}\text{C}$.

- 3.1.8 Explain why the boiling point of compound C is HIGHER than that of compound A and B by relating to the T_b and molecular of intermolecular forces. (2)

3.2 Substances D is added to water, whereas E is added to water. The diagram below shows the solution in the diagram below.



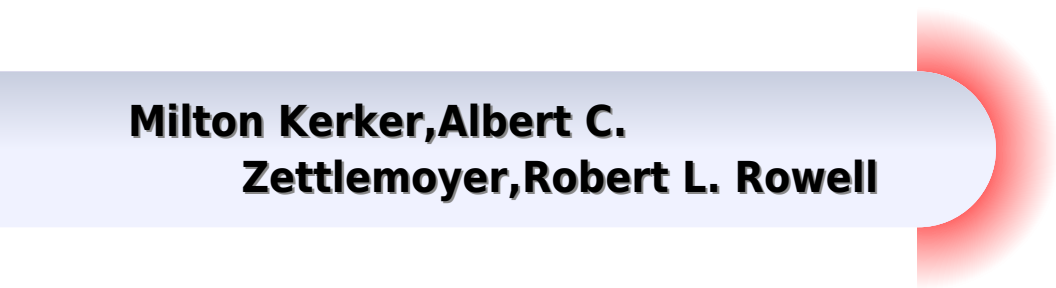
- 3.2.1 Is substance D or E, or both, ionic? Explain. (2)
- 3.2.2 Explain the process in QUESTION 3.1.1 about the solution in the intermolecular forces. (2)

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Physical Science Paper 2 Intermolecular Forces Investigation

**Milton Kerker, Albert C.
Zettlemoyer, Robert L. Rowell**



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