

Organic Compounds: carbon containing chains

Alkane- single bonds

Alkene- double bonds

Alkyne- triple bonds

Naming uses prefixes based on # carbons, then the root above

Depending on bond type

1 carbon = meth-

2 carbon = eth-

3 carbon = prop-

4 carbon = but-

5 carbon = pent-

6 carbon = hex-

7 carbon = hept-

8 carbon = oct-

9 carbon = non-

10 carbon = dec-

Monomers build up to be polymers

Nuclear Radiation: 1) elements above atomic # 83

2) atoms with uneven #p and #n ratio

Nuclear splitting Nuclear joining

Decay particles:

Alpha- α / ^4_2He , positive, most massive, like helium ion

Beta- β / $^-$ or β / $^+$, negative, neglig. Mass, like electron

Gamma- γ / $^0_0\gamma$, no charge, no mass, most energy, most penetrating

Half life time for half to decay

100 gram sample has half life of 10 days. How much remains after 30

days?

30/10 = 3 half lives: $100\text{g}/2 = 50\text{g}/2 = 25\text{g}/2 = 12.5\text{g}$, grams remain

Chemical Reactions: new substance created.

Chemical Equations

Reactant $\rightarrow$ Product

Synthesis: $A + B \rightarrow AB$

Decomposition: $AB \rightarrow A + B$

Single Displacement: $A + BC \rightarrow AC + B$

Double Displacement: $AB + CD \rightarrow AD + CB$

Balanced Chemical equations: Law of Conservation of Mass

Use coefficients- large whole numbers put front of substance

$x\text{H}_2 + \text{O}_2 \rightarrow x\text{H}_2\text{O}$

Coefficients also indicate # moles of each substance taking part in the reaction & mole ratio

Ex: 2 mol H_2 react with 1 mol O_2 to produce 2 mol H_2O

Molar mass: mass of 1 mole of a substance.

Add up masses from PT for each ATOM present in substance.

Ex: H_2O : [2 atoms of H @ 1g each] + [1 atom of oxygen @ 16g each] =

18 g/mol

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1. Understanding the eBook Physical Science Eoct Review Georgia Answer Key
 - The Rise of Digital Reading Physical Science Eoct Review Georgia Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Eoct Review Georgia Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Eoct Review Georgia Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Eoct Review Georgia Answer Key
 - Personalized Recommendations
 - Physical Science Eoct Review Georgia Answer Key User Reviews and Ratings
 - Physical Science Eoct Review Georgia Answer Key and Bestseller Lists
5. Accessing Physical Science Eoct Review Georgia Answer Key Free and Paid eBooks

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 - Physical Science Eoct Review Georgia Answer Key Budget-Friendly Options
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- ePub, PDF, MOBI, and More
 - Physical Science Eoct Review Georgia Answer Key Compatibility with Devices
 - Physical Science Eoct Review Georgia Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Physical Science Eoct Review Georgia Answer Key
 - Highlighting and Note-Taking Physical Science Eoct Review Georgia Answer Key
 - Interactive Elements Physical Science Eoct Review Georgia Answer Key
8. Staying Engaged with Physical Science Eoct Review Georgia Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Eoct Review Georgia Answer Key
9. Balancing eBooks and Physical Books Physical Science Eoct Review Georgia Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Eoct Review Georgia Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Eoct Review Georgia Answer Key
- Setting Reading Goals Physical Science Eoct Review Georgia Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Eoct Review Georgia Answer Key
- Fact-Checking eBook Content of Physical Science Eoct Review Georgia Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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