

CHEMISTRY

FORM FIVE

FIRST TERM – 03.01.2007 to 09.03.2007 (10 weeks)

19.03.2007 to 25.05.2007 (10 weeks)

THEME
LEARNING AREA: INTERACTION BETWEEN CHEMICALS
: 1. RATE OF REACTION

No. of week/ Weeks	Learning Objectives	Suggested Learning Activities	Learning Outcomes	Thinking Skills	Scientific Skills	Scientific Attitudes And Noble Values	Remarks/ Vocabulary
1	1.1 Analysing rate of reaction	Discusses : (a) the meaning of rate of reaction, (b) some examples of fast reactions, (c) some examples of slow reactions. Discusses to identify observable changes to reactants or products and its method of measurement in order to determine the rate of reaction. Carry out an activity involving a reaction between zinc and acid, and plot a graph to determine average rate of reaction and the rate of reaction at any given time. Carry out problem solving activities involving rates of reaction.	A student is able to: - state what rate of reaction is, - identify observable changes to reactants or products for determining rate of reaction, - determine average rate of reaction, - determine the rate of reaction at any given time from a graph, - solve numerical problems involving average rate of reaction, - solve numerical problem involving rate of reaction at any given time.	- generation of ideas - relating making hypothesis - making conclusion - problem solving	- Observing - Hypothesising - Communicating - Interpreting data	- Being honest and accurate in recording and validating data. - Being cooperative - Being systematic	reactant- bahan tindak balas product- hasil tindak balas rate of reaction- kadar tindak balas observable change- perubahan yang dapat diperhatikan

No. of week/ Weeks	Learning Objectives	Suggested Learning Activities	Learning Outcomes	Thinking Skills	Scientific Skills	Scientific Attitudes And Noble Values	Remarks/ Vocabulary
2	1.2 Synthesising factors affecting the rate of reaction	Discusses possible factors affecting the rate of reaction. Design and carry out activities to investigate factors affecting the rate of reaction, i.e. size of reactant, concentration, temperature and catalyst. Some suggested reactions: (a) a reaction between calcium carbonate, CaCO_3, and hydrochloric acid, HCl, (b) a reaction between sodium thiosulphate, $\text{Na}_2\text{S}_2\text{O}_3$, and sulphuric acid, H_2SO_4, (c) decomposition of hydrogen peroxide, H_2O_2, in the presence of a catalyst. Use computer simulations to investigate how the movement and collision of particles in a reaction are affected by temperature, size of reactant, pressure, concentration and catalyst. Collect and interpret data to explain factors affecting the rate of reaction in the following: (a) combustion of charcoal, (b) storing food in a	A student is able to: - design experiment to investigate factors affecting the rate of reaction, - give examples of reactions that are affected by size of reactant, concentration, temperature and catalyst, - explain how each factor affects the rate of reaction, - describe how factors affecting the rate of	- generation of ideas - relating making hypothesis - analyzing - contributing - making conclusion - predicting making inference - problem solving	- experimenting - predicting - interpreting data - inferring	- being honest and accurate in recording and validating data - being systematic - being cooperative - being confident and independent	catalyst - mampin decomposition- penguraian

Scheme Of Work Chemistry

David Kirk



Scheme Of Work Chemistry:

Regulations for the Training of Teachers for Elementary Schools .. Great Britain. Board of Education,1915
Parliamentary Papers Great Britain. Parliament. House of Commons,1915 Report Commonwealth Shipping Committee,1915 *Report of the Consultative Committee on Practical Work in Secondary Schools ...* Great Britain. Board of Education. Consultative Committee,1913 *Chemical Engineering and the Works Chemist* ,1922 Appendix to the Journal of the House of the Representatives New Zealand. Legislature. House of Representatives,1912 **Journal of the Society of Chemical Industry** Society of Chemical Industry (Great Britain),1909 **Rate Coefficients in Astrochemistry** T.J. Millar,D.A. Williams,2012-12-06 An atteJDpt has been made to cOll1PlJte the numbers of certain JI10lecules in interstellar space A search for the bands of CH O DR en and C2 would appear to be proIDising P Swings and L Rosenfeld Astrophysical Journal 86 483 1937 This may have been the first attempt at modelling interstellar chemistry As with models today the methods used lacked reliability but the speculation was impressive Mark Twain might well have said of this infant subject One gets such wholesale returns of conjecture out of such a trifling investment of fact The detection of unidentified lines around the period that Swings and Rosenfeld were writing provoked much interest but even the most optimistic speculator could hardly have imagined developments which would occur during the next 50 years By 1987 about 70 varieties of molecule had been identified in the interstellar and circumstellar regions They range in complexity from simple diatomics such as H₂ and CO to such species as ethanol C₂HeDH acetone CHs 2CO and the largest interstellar molecule detected so far cyano penta acetylene HC₁₁N The study of these molecules in astronomy has developed enormously especially over the last 20 years and is now codified in the new subject of astrochemistry That such a variety of chemical species should exist in tenuous regions of the Galaxy is fascinating **The Chemical World** ,1912 *Journal of the Society of Chemical Industry* ,1909

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