### Thermochemistry Review Questions

- The process of dissolving ammonium nitrate, NH<sub>4</sub>NO<sub>3</sub>, in water is an <u>endothermic</u> process. What is the sign of 'q'? If you were to add some ammonium nitrate to water in a flask, would you expect the flask to feel warm or cold?
- Nitric acid, a source of many nitrogen compounds, is produced from nitrogen dioxide. An old process for
  making nitrogen dioxide employed nitrogen and oxygen: N<sub>2(g)</sub> + 2O<sub>2(g)</sub> → 2NO<sub>2(g)</sub>. The reaction absorbs 56.2
  kJ of heat per 2 mol of NO<sub>2(g)</sub> produced. Is the reaction endothermic or exothermic? What is the value of q?
- Ammonia burns in the presence of a copper catalyst to form nitrogen gas: NH<sub>3(g)</sub> + 3O<sub>2(g)</sub> → 2N<sub>2(g)</sub> + 6H<sub>2</sub>O<sub>(g)</sub> with a ΔH = −1267 kJ/mol. What is the enthalpy change (ΔH) to burn 35.8 g of ammonia? [Ans: -2 663.45 kJ]
- Propane, C<sub>3</sub>H<sub>8</sub>, is a common fuel gas. Use the following to calculate the grams of propane you would need to provide 369 kJ of heat: C<sub>3</sub>H<sub>8(a)</sub> + 5O<sub>2(a)</sub> → 3CO<sub>2(a)</sub> + 4H<sub>2</sub>O<sub>(a)</sub> with a ΔH = −2043 kJ/mol. [Ans: 7.96 g]
- Ammonia burns in the presence of a platinum catalyst to form nitric oxide: 4NH<sub>3(g)</sub> + 5O<sub>2(g)</sub> → 4NO<sub>(g)</sub> + 6H<sub>2</sub>O<sub>(g)</sub>. What is the heat of reaction at constant pressure? Use the following thermochemical equations [Ans: -906.3 kJ]:

$$N_{2(g)} + O_{2(g)} \rightarrow 2NO_{(g)}$$
  $\Delta H = 180.6 \ kJ$   
 $N_{2(g)} + 3H_{2(g)} \rightarrow 2NH_{3(g)}$   $\Delta H = -91.8 \ kJ$   
 $2H_{2(g)} + O_{2(g)} \rightarrow 2H_{2}O_{(g)}$   $\Delta H = -483.7 \ kJ$ 

Hydrogen sulfide, H<sub>2</sub>S, is a poisonous gas with the odour of rotten eggs. The reaction for the formation of H<sub>2</sub>S from the element is: H<sub>2(g)</sub> + ½S<sub>B(rhombic)</sub> → H<sub>2</sub>S<sub>(g)</sub>. Use Hess' law to obtain the enthalpy change for this reaction from the following enthalpies [Ans: - 21 kJ]:

$$\begin{array}{ll} H_2S_{(g)} + \frac{3}{2}O_{2(g)} \to H_2O_{(g)} + SO_{2(g)} & \Delta H = -518 \, kJ \\ H_{2(g)} + \frac{1}{2}O_{2(g)} \to H_2O_{(g)} & \Delta H = -242 \, kJ \\ \frac{1}{8}S_{8(rhomble)} + O_{2(g)} \to SO_{2(g)} & \Delta H = -297 \, kJ \end{array}$$

- The cooling effect of alcohol on the skin is due to its evaporation. Calculate the heat of vaporization (ΔH) of ethanol, C<sub>2</sub>H<sub>5</sub>OH<sub>(I)</sub> → C<sub>2</sub>H<sub>5</sub>OH<sub>(g)</sub>. The standard enthalpy of C<sub>2</sub>H<sub>5</sub>OH<sub>(I)</sub> is 277.7 kJ/mol and that of C<sub>2</sub>H<sub>5</sub>OH<sub>(g)</sub> is 235.4 kJ/mol. [Ans: 42.3 kJ]
- Carbon disulfide is a colourless liquid. When pure, it is nearly odourless, but the commercial produce smells vile. Carbon disulfide is used in the manufacturer of rayon and cellophane. The liquid burns as follows: CS<sub>2(I)</sub> + 3O<sub>2(g)</sub> → CO<sub>2(g)</sub> + 2SO<sub>2(g)</sub>. Calculate the standard enthalpy change for this reaction using standard enthalpies of formation [Ans: -1076.6 kJ].
- The group IIA carbonates decompose when heated. For example: BaCO<sub>3(s)</sub> → BaO<sub>(s)</sub> + CO<sub>2(g)</sub>. Use the following enthalpies of formation and calculate the heat require to decompose 10 g of barium carbonate. The standard enthalpy of BaCO<sub>3</sub> is 1216.3 kJ/mol. BaO<sub>(s)</sub> is 942.3 kJ/mol and CO<sub>2(g)</sub> is 393.5 kJ/mol. [Ans: -6.09 kJ].

# **Sch4u Thermochemistry Questions**

**Arshad Iqbal** 

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