

9th Grade IGCSE Chemistry Semester Exam Review

Unit 1: Atomic Theory

- Describe the structure of an atom in terms of electrons and a nucleus containing protons and neutrons and state the relative charges and approximate relative of these sub-atomic particles
- Define *proton number* and *nucleon number*.
- Use proton number and the simple structure of atoms to explain the basis of the Periodic Table, with special reference to the elements of proton number 1 to 20.
- Define *isotopes*.
- Describe the build-up of electrons in 'shells' and understand the significance of the noble gas electronic structures and of valency electrons
- Demonstrate understanding of the terms atom and molecule.

Practice Questions

1.

What do the nuclei in a hydrogen molecule contain?

- A** electrons and neutrons
- B** electrons and protons
- C** neutrons only
- D** protons only

2.

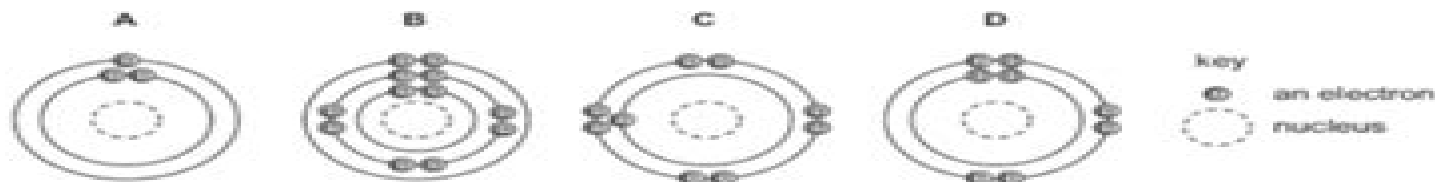
The proton number of helium is 2.

What information does this give about helium?

- A** Its atom has two electrons.
- B** Its atom is twice as heavy as a hydrogen atom.
- C** It is a Group II element.
- D** Its molecule has two atoms.

3.

Which diagram shows an atom in the same group of the Periodic Table as sodium?



4.

What is the electronic structure of an atom with a proton number 5 and a nucleon number 11?

- A** 1, 8, 2
- B** 2, 8, 1
- C** 2, 3
- D** 3, 2

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Roger Norris



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Table of Contents Review For Igcse Chemistry Semester Exam Xtremepapers

1. Understanding the eBook Review For Igcse Chemistry Semester Exam Xtremepapers
 - The Rise of Digital Reading Review For Igcse Chemistry Semester Exam Xtremepapers
 - Advantages of eBooks Over Traditional Books
2. Identifying Review For Igcse Chemistry Semester Exam Xtremepapers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review For Igcse Chemistry Semester Exam Xtremepapers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review For Igcse Chemistry Semester Exam Xtremepapers
 - Personalized Recommendations
 - Review For Igcse Chemistry Semester Exam Xtremepapers User Reviews and Ratings

- Review For Igcse Chemistry Semester Exam Xtremepapers and Bestseller Lists
- 5. Accessing Review For Igcse Chemistry Semester Exam Xtremepapers Free and Paid eBooks
 - Review For Igcse Chemistry Semester Exam Xtremepapers Public Domain eBooks
 - Review For Igcse Chemistry Semester Exam Xtremepapers eBook Subscription Services
 - Review For Igcse Chemistry Semester Exam Xtremepapers Budget-Friendly Options
- 6. Navigating Review For Igcse Chemistry Semester Exam Xtremepapers eBook Formats
 - ePub, PDF, MOBI, and More
 - Review For Igcse Chemistry Semester Exam Xtremepapers Compatibility with Devices
 - Review For Igcse Chemistry Semester Exam Xtremepapers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review For Igcse Chemistry Semester Exam Xtremepapers
 - Highlighting and Note-Taking Review For Igcse Chemistry Semester Exam Xtremepapers
 - Interactive Elements Review For Igcse Chemistry Semester Exam Xtremepapers
- 8. Staying Engaged with Review For Igcse Chemistry Semester Exam Xtremepapers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review For Igcse Chemistry Semester Exam Xtremepapers
- 9. Balancing eBooks and Physical Books Review For Igcse Chemistry Semester Exam Xtremepapers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review For Igcse Chemistry Semester Exam Xtremepapers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review For Igcse Chemistry Semester Exam Xtremepapers
 - Setting Reading Goals Review For Igcse Chemistry Semester Exam Xtremepapers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review For Igcse Chemistry Semester Exam Xtremepapers
 - Fact-Checking eBook Content of Review For Igcse Chemistry Semester Exam Xtremepapers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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