

CHAPTER 17. CHEMICAL EQUILIBRIUM

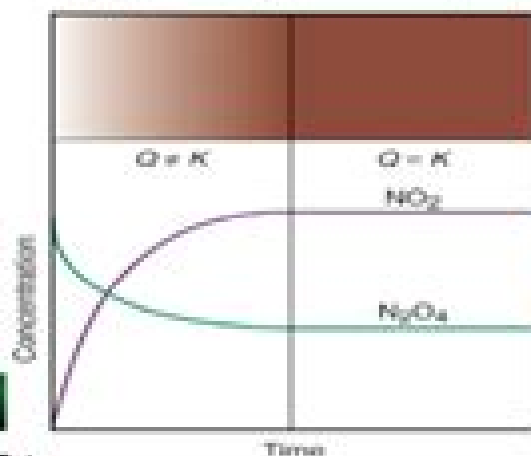
Section 17.1 Equilibrium State and Equilibrium Constant

Chemical reactions do NOT go to completion (100% products) - even those that look like they do. Reactions instead reach a point (called equilibrium) after which the amount of reactants and products no longer change with time. This is because *all reactions are reversible*.

Example: $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$ ($\rightleftharpoons$ means the reaction goes in both directions)
 (colorless) (brown)

--- Add some colorless $\text{N}_2\text{O}_4(\text{g})$ to an empty flask and then watch it with time - the gas will slowly go more and more brown as $\text{N}_2\text{O}_4(\text{g})$ converts to brown $\text{NO}_2(\text{g})$.

--- After a few minutes, no further color change occurs - "the reaction has reached equilibrium". The gas contains both N_2O_4 and NO_2 . Why doesn't it keep giving more NO_2 and going more brown? Because the rate at which N_2O_4 is giving more NO_2 is now the same as the rate of the reverse reaction (2NO_2 combining to give N_2O_4). Thus the $[\text{N}_2\text{O}_4]$ and $[\text{NO}_2]$ concs do not change (Figure 17.1). We can see this in a conc ([]) vs time (t) plot (Figure 17.3)



At any given time t:

Rate (forward) = $k_{\text{fwd}}[\text{N}_2\text{O}_4] = -\Delta[\text{N}_2\text{O}_4]/\Delta t$

Rate (reverse) = $k_{\text{rev}}[\text{NO}_2]^2 = -\Delta[\text{NO}_2]/\Delta t$

(Note: (i) assume reaction is elementary.

(ii) we'll define what Q is soon.)

After a certain time, the concs of all species do not change any more. The reaction has reached equilibrium. This is a "dynamic equilibrium", which means reactions are still occurring, but no overall change in the []s.

Review Chemical Equilibrium Section 2

Wenbin Ji



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