

UNIT 1 NOTES

Biological Chemistry Review

Why study Chemistry in Biology?

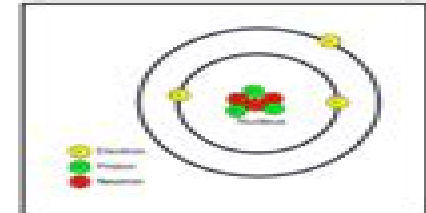
- The understanding of the basic chemistry of life is vital to a full comprehension of biology.
- Important atoms include:
 - Oxygen (O)
 - Carbon (C)
 - Hydrogen (H)
 - Nitrogen (N)
 - Phosphorous (P)
 - Potassium (K)
 - Sulphur (S)

Atoms:

Atom = smallest part of an element that displays the properties of that element

- Made up of 3 subatomic particles:
 - Protons (positively charged)
 - Neutrons (uncharged / neutral)
 - Electrons (negatively charged) – found outside of the nucleus

Found in the nucleus



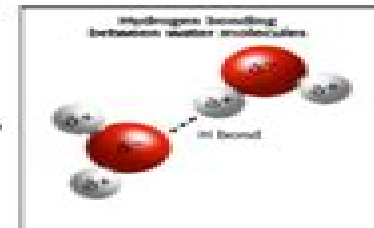
Ions:

Ions = atoms that have gained or lost electrons

- Signified by either (+) or (-)
- Ex. Na^+ , K^+ , Cl^- , Ca^{2+}

Types of bonding:

1. Ionic bonds:
 - Transfer of electrons from a metal (positive ion) to a non-metal (negative ion)
2. Covalent bonds:
 - Atoms share electrons between 2 non-metals
 - Can be single covalent bond (sharing 2 electrons), double covalent bond (sharing 4 electrons) or triple covalent bond (sharing 6 electrons)
 - Nonpolar covalent bonds
 - Electrons shared fairly equally between atoms
 - Polar covalent bonds
 - Unequal sharing of electrons
3. Hydrogen bonds:
 - When a covalently bonded H within a molecule (eg H_2O) is also attracted to another atom (usually O, N or F) of a **different** molecule.
 - Always represented by dotted lines



Inorganic Compounds:

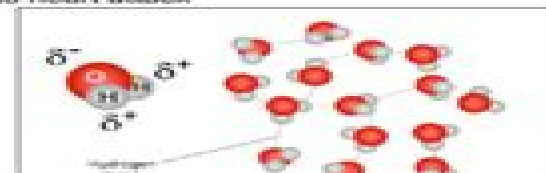
Inorganic Compounds	Organic Compounds
Usually metal + non-metal	Always contain carbon and hydrogen
Usually ionic bonding	Always covalent bonding
Associated with non-living	Associated with living things
Ex. Salts, water, acids, bases, buffers	Ex. Carbohydrates, lipids, proteins and nucleic acids

A. Salts:

- Needed for survival
- Too little → fatigue, exhaustion, cramps, nausea, thirst, decrease in blood pressure and even death
- Too much → dehydration of cells and hypertension; can lead to heart attack

B. Water:

- Structure of water:
- Biological importance is due to its chemical structure
- Water is a **polar** molecule
 - **How does polarity of water produce H-bonding?**
 - Oxygen is bigger & negatively charged
 - H is smaller & positively charged



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