

Classifying and Balancing Equations
Multiple Choice

Name _____
Date _____

Per _____

PRACTICE TEST

- _____ 1. A chemical reaction is a process in which
a. products change into reactants. c. substances can change state
b. the law of conservation of mass applies d. all of these
- _____ 2. During a chemical reaction,
a. new elements are produced c. atoms are destroyed
b. atoms are rearranged d. elements are destroyed
- _____ 3. An equation is balanced by
a. changing subscripts c. erasing elements as necessary
b. adding coefficients d. adding elements as necessary
- _____ 4. An atom's ability to undergo chemical reactions is determined by its
a. protons b. innermost electrons c. neutrons d. valence electrons
- _____ 5. What are the reactants in the following chemical equation:
$$\text{Zn} + \text{CuSO}_4 \longrightarrow \text{ZnSO}_4 + \text{Cu}$$

a. zinc and copper c. zinc and copper (II) sulfate
b. zinc sulfate and copper d. only zinc
- _____ 6. What are the products in the above equation?
a. zinc and copper c. zinc and copper (II) sulfate
b. zinc sulfate and copper d. only zinc

Short Answer

For questions 7-12, classify the reaction according to the type it is. Put that answer in the blank. Then add coefficients to balance the reaction when necessary.

- _____ 7. $\text{Zn} + \text{H}_2\text{SO}_4 \longrightarrow \text{ZnSO}_4 + \text{H}_2$
- _____ 8. $\text{H}_2\text{CO}_3 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- _____ 9. $\text{CaCO}_3 + \text{HCl} \longrightarrow \text{CaCl}_2 + \text{H}_2\text{CO}_3$
- _____ 10. $\text{AgNO}_3 + \text{Zn} \longrightarrow \text{Zn}(\text{NO}_3)_2 + \text{Ag}$
- _____ 11. $\text{C}_3\text{H}_8 + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- _____ 12. $\text{C}_2\text{H}_5\text{OH} + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$

Write a balanced equation for each of the following reactions:

13. Magnesium chloride is the product of a reaction between magnesium and chlorine.

14. Copper (II) hydroxide and potassium sulfate are produced when potassium hydroxide reacts with copper (II) sulfate.

Unit 6 Practice Test – Key

1. b
2. b
3. b
4. d
5. c
6. B
7. Single replacement 1,1,1,1
8. Decomposition 1,1,1
9. Double replacement 1,2,1,1-
 CaCO_3 is ppt (see Ref. Table)
10. Single replacement 2,1,1,2

11. Combustion 1, 5, 3, 4
12. Combustion 1,3,2,3
13. $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$ 1,1,1
14. $\text{KOH} + \text{CuSO}_4 \rightarrow \text{Cu}(\text{OH})_2 + \text{K}_2\text{SO}_4$ 2,1,1,1

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Table of Contents Review Chemical Reaction Unit Answers

1. Understanding the eBook Review Chemical Reaction Unit Answers
 - The Rise of Digital Reading Review Chemical Reaction Unit Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Chemical Reaction Unit Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Chemical Reaction Unit Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Chemical Reaction Unit Answers
 - Personalized Recommendations
 - Review Chemical Reaction Unit Answers User Reviews and Ratings
 - Review Chemical Reaction Unit Answers and Bestseller Lists

5. Accessing Review Chemical Reaction Unit Answers Free and Paid eBooks
 - Review Chemical Reaction Unit Answers Public Domain eBooks
 - Review Chemical Reaction Unit Answers eBook Subscription Services
 - Review Chemical Reaction Unit Answers Budget-Friendly Options
6. Navigating Review Chemical Reaction Unit Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Review Chemical Reaction Unit Answers Compatibility with Devices
 - Review Chemical Reaction Unit Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Chemical Reaction Unit Answers
 - Highlighting and Note-Taking Review Chemical Reaction Unit Answers
 - Interactive Elements Review Chemical Reaction Unit Answers
8. Staying Engaged with Review Chemical Reaction Unit Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Chemical Reaction Unit Answers
9. Balancing eBooks and Physical Books Review Chemical Reaction Unit Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Chemical Reaction Unit Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Review Chemical Reaction Unit Answers
 - Setting Reading Goals Review Chemical Reaction Unit Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Review Chemical Reaction Unit Answers
 - Fact-Checking eBook Content of Review Chemical Reaction Unit Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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