Chemical Equations and Reactions

SECTION 2

SHORT ANSWER Answer the following questions in the space provided.

1. Match the equation type on the left to its representation on the right,

___ synthesis

(a) AX + BY → AY + BX

___d ___ decomposition

(b) $A + BX \rightarrow AX + B$

b single-displacement

(c) A + B → AX (d) AX → A + X

- _a__ double-displacement
- In the reaction described by the equation 2Al(s) + 3Fe(NO₅)₂(aq) → 3Fe(s) + 2Al(NO₅)₂(aq), iron has been replaced by
 - (a) nitrate. (c) aluminum.
 - (b) water.

(d) nitrogen.

- Of the following chemical equations, the only reaction that is both synthesis and combustion is
 - (a) C(s) + O₂(g) → CO₂(g).
 - (b) $2C_4H_{10}(I) + 13O_2(g) \rightarrow 8CO_2(g) + 10H_2O(I)$.
 - (c) $6CO_2(g) + 6H_2O(g) \rightarrow C_6H_{12}O_6(aq) + 6O_2(g)$.
 - (d) $C_6H_{12}O_6(aq) + 6O_2(g) \rightarrow 6CO_2(aq) + 6H_2O(l)$.
- __b__ Of the following chemical equations, the only reaction that is both combustion and decomposition is
 - (a) S(s) + O₂(g) → SO₂(g).
 - (b) 2C₄H₁₀(I) + 13O₂(g) → 8CO₂(g) + 10H₂O(I).
 - (c) $2H_2O_2(I) \rightarrow 2H_2O(I) + O_2(g)$.
 - (d)
- Identify the products when the following substances decompose:
 - a. Its separate elements _____ a binary compound
 - b. Metal oxide + water _____ most metal hydroxides
 - c. Metal oxide + carbon dioxide a metal carbonate
 - d. Water + sulfur dioxide the acid H₂SO₃
- The complete combustion of a hydrocarbon in excess oxygen yields the products

Review Chemical Equations And Reactions Section 3

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Review Chemical Equations And Reactions Section 3:

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