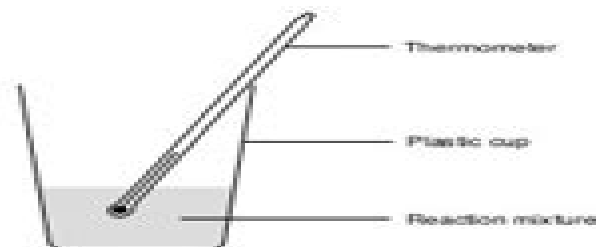


Exothermic or endothermic?

Introduction

Some reactions give out heat and others take in heat. In exothermic reactions the temperature goes up, in endothermic reactions the temperature goes down. In this experiment, various reactions are examined. Temperatures are measured to decide whether a particular reaction is exothermic or endothermic.



What to record

Complete the table

Reaction	Temperature before mixing/ $^{\circ}\text{C}$	Temperature after mixing/ $^{\circ}\text{C}$	Exothermic or endothermic
Sodium hydroxide solution + dilute hydrochloric acid			
Sodium hydrogen carbonate solution + citric acid			
copper(II) sulfate solution + magnesium powder			
Dilute sulfuric acid + magnesium ribbon			

What to do

1. Use the apparatus as shown.
2. Put 10 cm^3 of sodium hydroxide solution in the beaker, record the temperature then add 10 cm^3 of dilute hydrochloric acid, stirring with the thermometer. Record the maximum or minimum temperature.
3. Repeat the procedure for the following reactions: (a) sodium hydrogen carbonate solution and citric acid; (b) copper(II) sulfate solution and magnesium powder; and (c) dilute sulfuric acid and magnesium ribbon.

Safety

Wear eye protection. Some of the solutions are irritant.

Results For Acid And Alkali Reaction Exothermic

**Neena Sinha, R Rangarajan, R P
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