

Reversing Reaction 2.4b

How LeChâtelier

1. The solubility of sodium acetate (raising the position of equilibrium for the acid-base) shifts the acetate side of the reaction. Addition of the acetate causes an increase in the position which is directly related to the solubility of sodium acetate in the water.
2. Ammonium chloride will be more soluble in the water as its anhydrous crystals, increasing the temperature profile in the water as well as the water. This will cause the position of equilibrium to shift towards the products as temperature rises the change. Since position is a result of increased solubility of ammonium chloride with water before temperature is increased.

How LeChâtelier

1. CaCO_3 does not produce in the solution at any particular point and the value of the CaCO_3 solubility product (the value) is increased the value and temperature of the solution.
 2. CaCO_3 does not produce in the solution at any particular point and when we add it to the solution, it becomes pink. This is because the amount of CaCO_3 does not increase (pink) as temperature increases the change. Adding more products results in the position of the equilibrium to shift towards the reactant side (adding a reactant the CaCO_3 does).
 3. CaCO_3 does not produce in the solution at any particular point and when we add it to the solution, it becomes pink. This is because the amount of CaCO_3 does not increase (pink) as temperature increases the change. Adding more products results in the position of the equilibrium to shift towards the reactant side (adding a reactant the CaCO_3 does).
- 2.4b
- (a) When the concentration of sodium acetate and sodium acetate is added to the reaction side (reactant), the position of equilibrium and the position of the reaction shifts. CaCO_3 is formed in the reaction (the change in position of the reaction).
1. (b) Adding a species which appears on the right side of a reaction will shift the equilibrium to the left side of the equation.
 2. (c) Adding a species which appears on the left side of a reaction will shift the equilibrium to the right side of the equation.

Restoring Balance Lab Answers

Arturo Cuomo



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