

## Answers

$Q = mc\Delta T$ , where  $Q$  = heat energy,  $m$  = mass, and  $\Delta T$  = change in temp.  
Remember,  $\Delta T = (T_{\text{final}} - T_{\text{initial}})$ . Show all work and proper units.

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{1086.75}{15.75(175-25)} = 0.46 \text{ J/g}^\circ\text{C}$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$Q = mC(T_f - T_i) = 10.0\text{g} (0.90\text{J/g}^\circ\text{C})(55-22) = 297 \text{ J}$$

3. Calculate the specific heat capacity of a piece of wood if 1500.0 g of the wood absorbs 67,500 joules of heat, and its temperature changes from 32°C to 57°C.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{67500 \text{ J}}{(1500 \text{ g})(57-32)} = 1.8 \text{ J/g}^\circ\text{C}$$

4. 100.0 g of 4.0°C water is heated until its temperature is 37°C. Calculate the amount of heat energy needed to cause this rise in temperature.

$$Q = mC(T_f - T_i) = 100\text{g}(4.184\text{J/g}^\circ\text{C})(37 - 4) = 14000 \text{ J}$$

5. 25.0 g of mercury is heated from 25°C to 155°C, and absorbs 455 joules of heat in the process. Calculate the specific heat capacity of mercury.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{455 \text{ J}}{(25\text{g})(155-25)} = 0.14 \text{ J/g}^\circ\text{C}$$



# Physical Science If8767 Heat Calculations

**A Gutmann**



## **Physical Science If8767 Heat Calculations:**

**Heat, Thermodynamics and Radiation** Alauddin Khan, 2021-02-01 Document from the year 2020 in the subject Physics Thermodynamics grade 4 00 language English abstract The book consists of thirteen chapters to fulfill requirements of different kind of readers This volume takes into account the study of Thermometry Kinetic theory of gases the equation of state The change of state Transmission of heat First law of Thermodynamics Thermodynamic functions Second law of Thermodynamics Third law of Thermodynamics Maxwell's equation Clausius Clapeyron equation and Radiation Laws The volume contains illustrative examples of both the ideas and the methods The book is intended as a text book on Heat Thermodynamics and Radiation for undergraduate levels and also as a reference book for anyone who is interested in this field of enquiry The book is comprehensive enough to cover all the topics that are usually taught to upper undergraduate students of Physics Chemistry and Engineering This book will be useful to students and teachers in different universities around the world

**An Introduction to Heat Transfer** A. J. Ede, 2000      **Teaching Heat Transfer and Heat Exchange** Brian Melvin Evenson, 2004      *The New Heat Theorem* Walther Nernst, 1969      **The Heat Equation** D. V. Widder, 1976-01-22 The Heat Equation      **Heat-transfer Calculations by Finite Differences** George Merrick Dusenberre, 1961      **Calculations in Heat, Light and Sound ...** S. M. Niven, 1964      **Worked Problems in Heat, Thermodynamics, and Kinetic Theory for Physics Students** L. Pincherle, 1966 Worked Problems in Heat Thermodynamics and Kinetic Theory for Physics Students is a complementary to textbooks in physics This book is a collection of exercise problems that have been part of tutorial classes in heat and thermodynamics at the University of London This collection of exercise problems with answers that are fully worked out deals with various topics This book poses problems covering the definition of temperature such as calculating the assigned value of the temperature of boiling water under specific conditions This text also gives example of problems dealing with the fir

*Heat and Thermodynamics* Mark Waldo Zemansky, Richard H. Dittman, 1997 This respected text deals with large scale easily known thermal phenomena and then proceeds to small scale less accessible phenomena The wide range of mathematics used in Dittman and Zemansky's text simultaneously challenges students who have completed a course in impartial differential calculus without alienating those students who have only taken a calculus based general physics course Examples of calculations are presented shortly after important formulas are derived Students see the solutions of problems related to the formulas Actual thermodynamic experiments are explained in detail The student sees the applicability of abstract thermodynamic concepts and formulas to real situations

**The Theory of Heat** Thomas Preston, 1904

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