

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and which from it? (anion/cation) _____
which cation it gains and electron it loses from it? _____ (anion/cation)
2. How placed in the halogens group has _____ valence electrons, and number is _____ 1. number is same as. This is because of the _____ rules.
3. How _____ bond in the addition between oppositely charged ions.
How is chemical bond is _____
4. On a covalent bond, shared _____ electrons, represented by dot
bond bond, all are _____ electrons.
5. Different group Which bond from the shared completely is a covalent bond is to
_____ and which group is shared equally to the _____
How can predict if a compound is ionic, polar covalent, or nonpolar covalent by
looking at the _____
6. How ionic bond form up to _____ valence electrons. This bond is formed by the
particular bond that all have 8 valence electrons is called _____
7. Shared pair of electrons form _____ by between two or
more _____

B. Formulas for Ionic Compounds – Put in the right number

Write number in the box indicated by	Magnesium	Fluorine	Phosphorus	Chlorine

3. Write in the chemical formula for each of the following ions:
anion/cation

Write your own ionic compounds and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

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